

Name \_\_\_\_\_

**MULTIPLE CHOICE.** Choose the one alternative that best completes the statement or answers the question.

- The conjugate base of  $\text{HSO}_4^-$  is  
 A)  $\text{H}_2\text{SO}_4$       B)  $\text{SO}_4^{2-}$       C)  $\text{H}_3\text{SO}_4^+$       D)  $\text{HSO}_4^+$       E)  $\text{OH}^-$
- The conjugate acid of  $\text{HSO}_4^-$  is  
 A)  $\text{H}^+$       B)  $\text{H}_2\text{SO}_4$       C)  $\text{HSO}_4^+$       D)  $\text{SO}_4^{2-}$       E)  $\text{HSO}_3^+$
- The molar concentration of hydronium ion in pure water at  $25^\circ\text{C}$  is \_\_\_\_\_.  
 A) 7.00      B)  $1.0 \times 10^{-7}$       C) 1.00      D)  $1.0 \times 10^{-14}$       E) 0.00
- The molar concentration of hydroxide ion in pure water at  $25^\circ\text{C}$  is \_\_\_\_\_.  
 A) 1.00      B)  $1.0 \times 10^{-14}$       C)  $1.0 \times 10^{-7}$       D) 0.00      E) 7.00
- What is the pH of an aqueous solution at  $25^\circ\text{C}$  in which  $[\text{OH}^-]$  is 0.0025 M?  
 A) +2.60      B) -2.60      C) +11.40      D) -11.40      E) -2.25
- What is the pH of an aqueous solution at  $25^\circ$  that contains  $3.98 \times 10^{-9}$  M hydronium ion?  
 A) 7.000      B) 9.000      C) 8.400      D) 5.600      E) 3.980
- What is the concentration (in M) of hydronium ions in a solution at  $25^\circ\text{C}$  with  $\text{pH} = 4.282$ ?  
 A) 4.28      B)  $1.92 \times 10^{-10}$       C)  $1.66 \times 10^{-4}$       D) 9.71      E)  $5.22 \times 10^{-5}$
- Which solution below has the highest concentration of hydroxide ions?  
 A)  $\text{pH} = 3.21$       B)  $\text{pH} = 9.82$       C)  $\text{pH} = 7.93$       D)  $\text{pH} = 12.59$       E)  $\text{pH} = 7.00$
- What is the pH of a 0.015-M aqueous solution of barium hydroxide?  
 A) 12.18      B) 1.52      C) 12.48      D) 1.82      E) 10.35
- Sodium hydroxide is a strong base. This means that \_\_\_\_\_.  
 A) aqueous solutions of NaOH contain equal concentrations of  $\text{H}^+$  (aq) and  $\text{OH}^-$  (aq)  
 B) NaOH does not dissociate at all when it is dissolved in water  
 C) NaOH dissociates completely to  $\text{Na}^+$ (aq) and  $\text{OH}^-$ (aq) when it dissolves in water  
 D) NaOH cannot be neutralized by a weak acid  
 E) NaOH cannot be neutralized by ordinary means
- Of the following acids, \_\_\_\_\_ is not a strong acid.  
 A)  $\text{HNO}_3$       B)  $\text{HCl}$       C)  $\text{HNO}_2$       D)  $\text{HClO}_4$       E)  $\text{H}_2\text{SO}_4$
- Of the following, \_\_\_\_\_ is a weak acid.  
 A)  $\text{HBr}$       B)  $\text{HF}$       C)  $\text{HClO}_4$       D)  $\text{HCl}$       E)  $\text{HNO}_3$

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and Assessing Understanding Section VIII The book's final section IX is devoted to Professional Issues facing college and university faculty who choose to adopt active learning in their courses The common feature underlying all of the strategies described in this book is their emphasis on actively engaging students who seek to make sense of natural objects and events Many of the strategies we highlight emerge from a constructivist view of learning that has gained widespread acceptance in recent years In this view learners make sense of the world by forging connections between new ideas and those that are part of their existing knowledge base For most students that knowledge base is riddled with a host of naïve notions misconceptions and alternative conceptions they have acquired throughout their lives To a considerable extent the job of the teacher is to coax out these ideas to help students understand how their ideas differ from the scientifically accepted view to assist as students restructure and reconcile their newly acquired knowledge and to provide opportunities for students to evaluate what they have learned and apply it in novel circumstances Clearly this prescription demands far more than most college and university scientists have been prepared for

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