

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
2. Calculate pH and pOH for weak acids and bases, and calculate K_a and K_b for weak acids and bases.
3. Calculate the concentration of weak acids and bases.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-----------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the chemical formula for each of the following acids, bases, and salts, and calculate the K_a and K_b for each of the following acids, bases, and salts.

1. K_a of $HClO_4$ is 1.0×10^{-1} (Equation 19.1)

2. K_b of OH^- is 1.0×10^{-14} (Equation 19.2)

3. K_a of $HClO_4$ is 1.0×10^{-1} (Equation 19.1)

4. K_b of OH^- is 1.0×10^{-14} (Equation 19.2)

5. K_a of $HClO_4$ is 1.0×10^{-1} (Equation 19.1)

6. K_b of OH^- is 1.0×10^{-14} (Equation 19.2)

7. K_a of $HClO_4$ is 1.0×10^{-1} (Equation 19.1)

8. K_b of OH^- is 1.0×10^{-14} (Equation 19.2)

9. K_a of $HClO_4$ is 1.0×10^{-1} (Equation 19.1)

10. K_b of OH^- is 1.0×10^{-14} (Equation 19.2)

11. K_a of $HClO_4$ is 1.0×10^{-1} (Equation 19.1)

1. K_a of $HClO_4$ is 1.0×10^{-1} (Equation 19.1)
2. K_b of OH^- is 1.0×10^{-14} (Equation 19.2)
3. K_a of $HClO_4$ is 1.0×10^{-1} (Equation 19.1)
4. K_b of OH^- is 1.0×10^{-14} (Equation 19.2)
5. K_a of $HClO_4$ is 1.0×10^{-1} (Equation 19.1)
6. K_b of OH^- is 1.0×10^{-14} (Equation 19.2)
7. K_a of $HClO_4$ is 1.0×10^{-1} (Equation 19.1)
8. K_b of OH^- is 1.0×10^{-14} (Equation 19.2)
9. K_a of $HClO_4$ is 1.0×10^{-1} (Equation 19.1)
10. K_b of OH^- is 1.0×10^{-14} (Equation 19.2)
11. K_a of $HClO_4$ is 1.0×10^{-1} (Equation 19.1)

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