

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-----------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the appropriate formula for each of the following acids, bases, and salts, showing dissociation reactions, and calculate the pH of each solution.

1. Acetic acid, CH_3COOH , 0.10 M
2. Hydrofluoric acid, HF , 0.10 M
3. Hydrochloric acid, HCl , 0.10 M
4. Hydrobromic acid, HBr , 0.10 M
5. Hydroiodic acid, HI , 0.10 M
6. Nitric acid, HNO_3 , 0.10 M
7. Sulfuric acid, H_2SO_4 , 0.10 M
8. Phosphoric acid, H_3PO_4 , 0.10 M
9. Acetic acid, CH_3COOH , 0.10 M
10. Hydrofluoric acid, HF , 0.10 M
11. Hydrochloric acid, HCl , 0.10 M
12. Hydrobromic acid, HBr , 0.10 M
13. Hydroiodic acid, HI , 0.10 M
14. Nitric acid, HNO_3 , 0.10 M
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