

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization | 5. K_a | 6. K_b |
| 7. ionization constant | 8. pH | |

Key Equation

- | | |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

Part A Completion

Write the requested unit or numerical information of the values provided that are indicated in parentheses. Significant figures are indicated with the appropriate symbol.

1. What substance is _____, a weak hydrogen ion donor?
2. _____ is the concentration of those species present in water that equal to _____ units.
3. _____ is the concentration of those species present in _____ solutions.
4. _____ is the concentration of those species present in _____ solutions.
5. _____ is the concentration of those species present in _____ solutions.
6. _____ is the concentration of those species present in _____ solutions.
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19. _____ is the concentration of those species present in _____ solutions.
20. _____ is the concentration of those species present in _____ solutions.

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Eileen Ramsden



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