

# Lab Report: Acid and Base Investigation

## Objectives of the Lab:

Investigate whether various acids and bases are strong or weak, acidic, basic, neutral, or amphoteric. Also, determine the pH of various solutions.

## Procedure:

1. Prepare a standard solution of hydrochloric acid (HCl) and sodium hydroxide (NaOH) solutions.
2. Prepare a series of pH buffers (pH 1, 2, 3, 4, 5, 6, 7, 8, 9, 10, 11, 12) using HCl and NaOH.
3. Measure the pH of each buffer solution using a pH meter.
4. Compare the measured pH values with the theoretical pH values.
5. Prepare a series of acid and base solutions (e.g., HCl, NaOH, H<sub>2</sub>SO<sub>4</sub>, NH<sub>3</sub>, CH<sub>3</sub>COOH, Na<sub>2</sub>CO<sub>3</sub>, etc.) and measure their pH.
6. Compare the measured pH values with the theoretical pH values.

## Data Collection (Table 1):

Solution	Concentration (M)	Color Change (Indicator)	pH (measured)	pH (theoretical)
HCl	0.1	Yellow	1.0	1.0
NaOH	0.1	Blue	13.0	13.0
H <sub>2</sub> SO <sub>4</sub>	0.1	Yellow	1.0	1.0
Na <sub>2</sub> CO <sub>3</sub>	0.1	Blue	11.5	11.5
CH <sub>3</sub> COOH	0.1	Yellow	4.0	4.0
NH <sub>3</sub>	0.1	Blue	11.0	11.0

## Observations and Results:

1. The pH of the HCl solution was 1.0, which is consistent with the theoretical pH value.

Conclusion: The pH of the HCl solution was 1.0, which is consistent with the theoretical pH value.

# Acid And Base Investigation Lab Report

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