

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                 |                                      |                 |
|-----------------|--------------------------------------|-----------------|
| acid            | acid dissociation constant ( $K_a$ ) | acid ionization |
| acidic solution | acidic solution                      | acidic solution |
| acidic solution | acidic solution                      | acidic solution |

## Key Equations

- |                                 |                                  |
|---------------------------------|----------------------------------|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ | $K_b = \frac{[OH^-][HB^+]}{[B]}$ |
| $pH = -\log[H^+]$               | $pOH = -\log[OH^-]$              |

## Part A Completion

Write the appropriate number in the blank space for each of the following questions. (If the answer is a number, write the number in the blank space.)

1. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is \_\_\_\_\_.
2. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is \_\_\_\_\_.
3. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is \_\_\_\_\_.
4. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is \_\_\_\_\_.
5. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is \_\_\_\_\_.
6. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is \_\_\_\_\_.
7. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is \_\_\_\_\_.
8. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is \_\_\_\_\_.
9. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is \_\_\_\_\_.
10. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is \_\_\_\_\_.

1. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is \_\_\_\_\_.
2. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is \_\_\_\_\_.
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**University of Illinois at Urbana-  
Champaign**



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