

Limiting Reagents

In many reactions, one reactant becomes entirely consumed and the other reactant, or reactants, is left over. The reactant that becomes consumed is called the **limiting reactant** or **limiting reagent** or **limiting factor**. It is called the limiting reactant because when it is all gone no more product can be created – the reactant limits the amount of product formed. The reactants remaining are called the **excess reagents**.

Analogy: The Classic Grilled Cheese Sandwich

To make a grilled cheese sandwich you need two pieces of bread and one piece of cheese. The reaction would look something like this:



So, if we had 6 pieces of bread and 2 pieces of cheese, how many sandwiches could we make? We would be able to make 2 sandwiches because we only have 2 slices of cheese. We would end up with 2 slices of bread left over. This would mean that the cheese was our limiting reagent and the bread would be our excess reagent.

Example 1:

Sodium chloride can be prepared by the reaction of sodium metal with chlorine gas.



What will occur when 6.70 mol of Na reacts with 3.20 mole of Cl₂?

- What is the limiting reagent?
- How many moles of NaCl are produced?
- How much of the excess reagent remains unreacted?

Solution:

- Start with the known amount of one of the reactants. Use the mole ratio from the balanced equation to calculate the required amount of the other reactant.

It doesn't matter if you choose Na or Cl₂. We will choose Na.



$$\underbrace{6.7 \text{ mol Na}}_{\text{given amount of sodium from the question}} \times \frac{1 \text{ mol Cl}_2}{2 \text{ mol Na}} = \underline{3.35 \text{ mol Cl}_2}$$

required amount of Cl₂ if we used up all 6.70 mol Na from the balanced equation

* This calculation shows that 3.35 mol of Cl₂ is needed to react with all 6.70 mol of Na. We only have 3.20 mol C available. ∴ Cl₂ is the Limiting Reagent.

Answers For Quick Lab Limiting Reagents Section 12 3

James Joseph Gallagher

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