

## Limiting Reagents

In many reactions, one reactant becomes entirely consumed and the other reactant, or reactants, is left over. The reactant that becomes consumed is called the **limiting reactant** or **limiting reagent** or **limiting factor**. It is called the limiting reactant because when it is all gone no more product can be created – the reactant limits the amount of product formed. The reactants remaining are called the **excess reagents**.

*Analogy: The Classic Grilled Cheese Sandwich*

To make a grilled cheese sandwich you need two pieces of bread and one piece of cheese. The reaction would look something like this:



So, if we had 6 pieces of bread and 2 pieces of cheese, how many sandwiches could we make? We would be able to make 2 sandwiches because we only have 2 slices of cheese. We would end up with 2 slices of bread left over. This would mean that the cheese was our limiting reagent and the bread would be our excess reagent.

*Example 1:*

Sodium chloride can be prepared by the reaction of sodium metal with chlorine gas.



What will occur when **6.70 mol of Na** reacts with **3.20 mole of  $Cl_2$** ?

- What is the limiting reagent?
- How many moles of NaCl are produced?
- How much of the excess reagent remains unreacted?

*Solution:*

- Start with the known amount of one of the reactants. Use the mole ratio from the balanced equation to calculate the required amount of the other reactant.

It doesn't matter if you choose Na or  $Cl_2$ . We will choose Na.



$$\underbrace{6.7 \text{ mol Na}}_{\text{given amount of sodium from the question}} \times \underbrace{\frac{1 \text{ mol } Cl_2}{2 \text{ mol Na}}}_{\text{mole ratio (using coefficients from the balanced equation)}} = \underbrace{3.35 \text{ mol } Cl_2}_{\text{required amount of } Cl_2 \text{ if we used up all 6.70 mol Na}}$$

\* This calculation shows that 3.35 mol of  $Cl_2$  is needed to react with all 6.70 mol of Na. We only have **3.2 mol  $Cl_2$**  available.  $\therefore Cl_2$  is the Limiting Reactant.

# Answers For Quick Lab Limiting Reagents Section 12 3

**James Joseph Gallagher**



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## **Table of Contents Answers For Quick Lab Limiting Reagents Section 12 3**

1. Understanding the eBook Answers For Quick Lab Limiting Reagents Section 12 3
  - The Rise of Digital Reading Answers For Quick Lab Limiting Reagents Section 12 3
  - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Quick Lab Limiting Reagents Section 12 3
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answers For Quick Lab Limiting Reagents Section 12 3
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Quick Lab Limiting Reagents Section 12 3
  - Personalized Recommendations
  - Answers For Quick Lab Limiting Reagents Section 12 3 User Reviews and Ratings
  - Answers For Quick Lab Limiting Reagents Section 12 3 and Bestseller Lists

5. Accessing Answers For Quick Lab Limiting Reagents Section 12 3 Free and Paid eBooks
  - Answers For Quick Lab Limiting Reagents Section 12 3 Public Domain eBooks
  - Answers For Quick Lab Limiting Reagents Section 12 3 eBook Subscription Services
  - Answers For Quick Lab Limiting Reagents Section 12 3 Budget-Friendly Options
6. Navigating Answers For Quick Lab Limiting Reagents Section 12 3 eBook Formats
  - ePub, PDF, MOBI, and More
  - Answers For Quick Lab Limiting Reagents Section 12 3 Compatibility with Devices
  - Answers For Quick Lab Limiting Reagents Section 12 3 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Answers For Quick Lab Limiting Reagents Section 12 3
  - Highlighting and Note-Taking Answers For Quick Lab Limiting Reagents Section 12 3
  - Interactive Elements Answers For Quick Lab Limiting Reagents Section 12 3
8. Staying Engaged with Answers For Quick Lab Limiting Reagents Section 12 3
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answers For Quick Lab Limiting Reagents Section 12 3
9. Balancing eBooks and Physical Books Answers For Quick Lab Limiting Reagents Section 12 3
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answers For Quick Lab Limiting Reagents Section 12 3
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Answers For Quick Lab Limiting Reagents Section 12 3
  - Setting Reading Goals Answers For Quick Lab Limiting Reagents Section 12 3
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers For Quick Lab Limiting Reagents Section 12 3
  - Fact-Checking eBook Content of Answers For Quick Lab Limiting Reagents Section 12 3
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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