

# Chemistry Study Guide

Periodic Table of the Elements																		
H	Li	Be																
Na	Mg																	
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr	
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe	
Cs	Ba	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn		
Fr	Ra	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Uut	Fl	Uup	Lv	Uus	Uuo		
		La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu		
		Ac	Th	Pa	U	Ng	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr		

## Objectives:

To understand the structure of an atom, and the nature + properties of subatomic particles.



## Matter:

Anything that has mass and occupies space.

Examples: air, gases, liquids, solidz!

## Atoms:

Atoms are the smallest, indivisible particle.

- Outer part → shell/orbit/energy level → carries the electronz (e-)
- Inner part → nucleus → houses the protonz (p+) and neutronz (n)
- Outer & Inner part → together these are the subatomic particles.

## Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p <sup>+</sup>	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e <sup>-</sup>	-0	-1	Orbit/Shell

# 100 Question Study Guide For Chemistry

**Hongru Du**

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**General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing,1999      **Study Guide and Solutions Manual for Organic Chemistry** Susan McMurry,1992 John McMurry's best selling text presents organic chemistry in a new edition that is up to date beautifully written visually striking and pedagogically sound. Described by many of its users as an eminently teachable text McMurry sets the standard in the field. The writing style has received almost universal acclaim from its users. McMurry introduces new concepts only as needed and immediately illustrates them with concrete examples. And wherever possible he ties material together with brief reviews overviews and reaction summaries. The result is a text that helps students mentally organize the material a text that helps them understand concepts not just memorize facts and a text that helps them make sense of the voluminous amount of material they encounter in the study of organic chemistry. McMurry uses a simple but important polar reaction the addition of HBr to an alkene as the lead off reaction to illustrate the general principles of organic reactions. Users of former editions found this an excellent choice because of its relative simplicity no prior knowledge of chirality or kinetics is required and its importance as a polar reaction on a common functional group that offers students the key to understanding hundreds of thousands of ionic reactions. By selecting this particular model McMurry is able to offer an unusually early presentation of organic reactions

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