

Acids and bases review sheet Answer key

1. List the differences between an acid and a base. Your description should include such things as whether both are electrolytes, which contains a hydrogen ion and which includes a hydroxide ion, and you should also include two examples of an acid and two examples of a base that are present in your household.

Acid	base
<p>Conducts electricity</p> <p>Contains hydrogen ion</p> <p>Proton donor</p> <p>Turns litmus paper red</p> <p>React with most metals</p> <p>Releases hydronium</p> <p>Ex: lemon</p> <p>Vinegar</p> <p>Battery acid (sulfuric acid)</p>	<p>conducts electricity</p> <p>contains hydroxide ion</p> <p>proton acceptor</p> <p>turns litmus paper blue</p> <p>releases hydroxide ion</p> <p>Ex: Windex</p> <p>bleach</p> <p>soap</p>

2. Listed below are some of the properties of acids and bases. Fill in the blanks with the appropriate word, acids or bases:

- a) A produce hydrogen ions (H^+) in solution
- b) B have a bitter taste
- c) B produce hydroxide ions (OH^-) ions
- d) A react with bases to form salts
- e) A have a sour taste
- f) B have a slippery, soapy feel
- g) B react with acids to form salts
- h) A react with many metals

3. Write an equation for the dissociation of the following acids in water. Follow example.

a) $\text{HClO}_4 \rightarrow \text{H}^+ + \text{ClO}_4^-$ d) $\text{H}_2\text{S} \rightarrow 2\text{H}^+ + \text{S}^{2-}$
 b) $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$ e) $\text{HCl} \rightarrow \text{H}^+ + \text{Cl}^-$
 c) $\text{HC}_2\text{H}_5\text{O}_2 \rightarrow \text{H}^+ + \text{C}_2\text{H}_5\text{O}_2^-$ f) $\text{HNO}_3 \rightarrow \text{H}^+ + \text{NO}_3^-$

4. Write an equation for the dissociation (ionization) of the following bases in water: (follow example).

a) $Mg(OH)_2 \rightarrow Mg^{2+} + 2OH^-$
b) $NaOH \rightarrow Na^+ + OH^-$
c) $Ba(OH)_2 \rightarrow Ba^{2+} + 2(OH^-)$
d) $KOH \rightarrow K^+ + OH^-$
e) $Mg(OH)_2 \rightarrow Mg^{2+} + 2(OH^-)$
f) $LiOH \rightarrow Li^+ + OH^-$

Q: What type of reaction is a neutralization reaction? Double replacement

B. Write the net equation for neutralization:



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