

## Experiment 11

## **Data and Calculations: The Atomic Spectrum of Hydrogen**

#### A. Calculation of the Energy Levels of the Hydrogen Atom

Resistances are also calculated from Equations 6 and 7 for the 10 lowest energy states.

Table 11.2

## B. Calculation of the Wavelengths of the Lines in the Hydrogen Atomic Spectrum

In the upper half of each box, write  $\Delta E$ , the difference in energy in kJ/mole between  $E_{\infty}$  and  $E_{\text{rec}}$ . In the lower half of the box, write the  $\Delta$  in nm associated with that value of  $\Delta E$ .

Table 11.3

	6	5	4	3	2	1
1	12.75. 4 KTImo1	12.59. 5. 6 KTImo1	12.30. 0. 04 KTImo1	11.6. 1. 25 KTImo1	9.7. 0. 05 KTImo1	
	93.78 nm	94.91 nm	91.15 nm	101.51 nm	121.36 nm	
2	2.91. 57 KTImo1	2.75. 53 KTImo1	2.60. 01 KTImo1	1.71. 73 KTImo1		
	410.23 nm	434.16 nm	406.26 nm	656.46 nm		
3	109. 34 KTImo1	93. 3 KTImo1	63. 73 KTImo1			
	1694.6% nm	1251.01 nm	13. 15. 61 nm			
4	15. 56 KTImo1	29. 52 KTImo1				
	2.675. 70 nm	4.052. 46 nm				
5	16. 04 KTImo1					
	74656. 04 nm					

$$\Delta E = E_{\text{fin}} - E_{\text{ini}}$$

$$\lambda, [\text{nm}] = \frac{1.19627 \times 10^9}{\Delta E}$$

(continued on following page)

# Atomic Spectrum Of Hydrogen Lab Answer Key

**Judd E. Hollander**

**Atomic Spectrum Of Hydrogen Lab Answer Key:**

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