

Atomic Spectra: Answer Sheet

Name Madi Johnson

Section Number

035

Substance in the lamp	Color seen with eye	Number of lines observed using glasses (estimate)	Colors of lines in spectrum
Hydrogen	violet	3	Purple, blue, red
Neon	red	20	green yellow, orange, red
Mercury	indigo	4	yellow, green, blue, purple

Substance in flame	Chemical Formula	Flame Color
Lithium chloride	<chem>LiCl</chem>	pink, red
Sodium chloride	<chem>NaCl</chem>	orange
Potassium chloride	<chem>KCl</chem>	blue, orange
Barium chloride	<chem>BaCl2</chem>	yellow
Copper (II) chloride	<chem>CuCl2</chem>	indigo / green, blue
Strontium chloride	<chem>SrCl2</chem>	red, orange

QUESTIONS

1. A fluorescent light bulb contains hot gaseous mercury atoms. An incandescent bulb contains a thin piece of metal called a filament. It is heated until it glows, a phenomenon known as blackbody radiation. Based on what you know about atomic spectra, which light source will have a continuous spectrum, and which will have a noncontinuous spectrum? Explain.




The fluorescent light bulb will have a non continuous spectrum because as we viewed before mercury had a spectrum broken into lines. The incandescent light bulb will have a continuous spectrum because it is created by thermal energy.

2. Each element has its own distinct emission color bands. Why are the emission spectra of no two elements the same?

The emission spectra of no two elements are the same because each element have a different number of electrons and valence shells.

Atomic Spectra Lab Answer Key

**S. K. Joshi, B. D. Shrivastava, A. P.
Deshpande**



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