

Name: Maddi Johnson
DATA

Atomic Spectra: Answer Sheet

Section Number: 035

Substance in the lamp	Color seen with eye	Number of lines observed using glasses (estimate)	Colors of lines in spectrum
Hydrogen	violet	3	purple, olive, red
Neon	red	20	green, yellow, orange, red
Mercury	indigo	4	yellow, green, olive, purple

Substance in flame	Chemical Formula	Flame Color
Lithium chloride	LiCl	pink, red
Sodium chloride	NaCl	orange
Potassium chloride	KCl	pink, orange
Banum chloride	BaCl ₂	yellow
Copper (II) chloride	CuCl ₂	indigo, green, olive
Strontium chloride	SrCl ₂	red, orange

QUESTIONS

1. A fluorescent light bulb contains hot gaseous mercury atoms. An incandescent bulb contains a thin piece of metal called a filament. It is heated until it glows, a phenomenon known as blackbody radiation. Based on what you know about atomic spectra, which light source will have a continuous spectrum, and which will have a noncontinuous spectrum? Explain.



2. Each element has its own distinct emission color bands. Why are the emission spectra of no two elements the same?

The fluorescent light bulb will have a non continuous spectrum because as we viewed before mercury had a spectrum broken into lines. The incandescent light bulb will have a continuous spectrum because it is created by thermal energy.

The emission spectra of no two elements are the same because each element have a different number of electrons and valence shells.

Atomic Spectra Lab Answer Key

**S. K. Joshi, B. D. Shrivastava, A. P.
Deshpande**

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