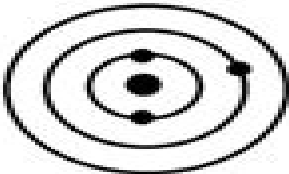
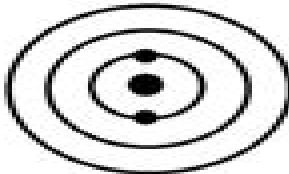
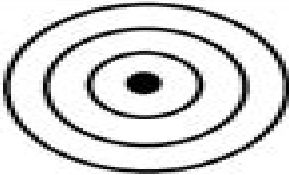
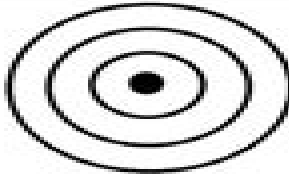
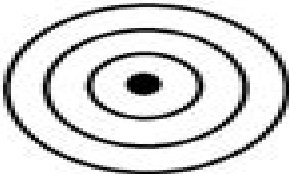
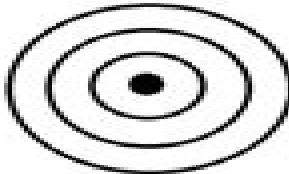


Cations	Anions
Have a positive charge because they _____ electrons	Have a negative charge because they _____ electrons
<p style="text-align: center;"><u>Ion symbol</u></p> <p>To write the ion symbol, you must write the element symbol with the charge written on the top right. Example: Ca^{2+}, Zn^{2+}, Ag^{1+}</p>	

For each box, use the Bohr diagram to show the number of electrons in a standard atom. Then, you will need to determine if the atom would lose or gain electrons and draw the Bohr diagram to show the number of electrons in the ion. The first one has been done for you.

Lithium	Lithium Atom	Lithium Ion	Lithium Ion
Group <u>1</u> , Period <u>2</u>		Atomic #: 3	
Atomic #: 3		Electrons: 2	
Electrons: 3		Cation/Anion: Cation	
Valence Electrons: 1		Ion Symbol: Li^{1+}	

Beryllium	Beryllium Atom	Beryllium Ion	Beryllium Ion
Group _____, Period _____		Atomic #:	
Atomic #:		Electrons:	
Electrons:		Cation/Anion:	
Valence Electrons:		Ion Symbol:	

Nitrogen	Nitrogen Atom	Nitrogen Ion	Nitrogen Ion
Group _____, Period _____		Atomic #:	
Atomic #:		Electrons:	
Electrons:		Cation/Anion:	
Valence Electrons:		Ion Symbol:	

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JA Banks

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the JA Banks text.

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