

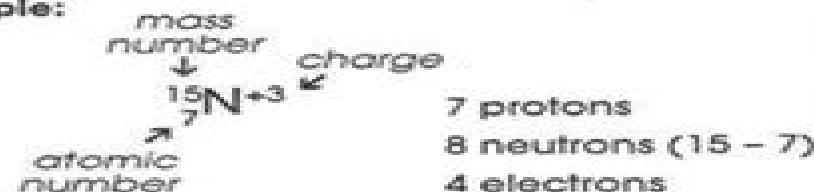
# ATOMIC STRUCTURE

Name \_\_\_\_\_

An atom is made up of protons and neutrons (both found in the nucleus) and electrons (in the surrounding electron cloud). The atomic number is equal to the number of protons. The mass number is equal to the number of protons plus neutrons. In a neutral atom, the number of protons equals the number of electrons. The charge on an ion indicates an imbalance between protons and electrons. Too many electrons produces a negative charge, too few, a positive charge.

This structure can be written as part of a chemical symbol.

Example:



Complete the following chart.

| Element/<br>Ion       | Atomic<br>Number | Atomic Mass | Mass<br>Number | Protons | Neutrons | Electrons |
|-----------------------|------------------|-------------|----------------|---------|----------|-----------|
| H                     |                  |             |                |         |          |           |
| H <sup>+</sup>        |                  |             |                |         |          |           |
| $^{12}\text{C}$       | 6                |             |                |         |          |           |
| $^7\text{Li}^+$       | 3                |             |                |         |          |           |
| $^{35}\text{Cl}^-$    | 17               |             |                |         |          |           |
| $^{39}\text{K}$       | 19               |             |                |         |          |           |
| $^{24}\text{Mg}^{2+}$ | 12               |             |                |         |          |           |
| As <sup>3-</sup>      |                  |             |                |         |          |           |
| Ag                    |                  |             |                |         |          |           |
| Ag <sup>+</sup>       |                  |             |                |         |          |           |
| S <sup>2-</sup>       |                  |             |                |         |          |           |
| U                     |                  |             |                |         |          |           |

# Atomic Structure Worksheet Chemistry If8766

**Brian Stevens**

## Atomic Structure Worksheet Chemistry If8766:

Chemistry & Atomic Structure John David Main Smith,1924    Atomic and Molecular Structure Joel Chaffee,2011-01-15

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## **ATOMIC STRUCTURE**

**High School Chemistry Curriculum** Bud Morgan, 2025-07-09 Welcome to our comprehensive guide to Atomic Structure designed specifically for high school students studying chemistry. In this book we will explore the fascinating world of atoms, the basic building blocks of matter and delve into the complexities of their structure. Chapter 1 provides a solid foundation by introducing the concept of atoms and their historical significance in the development of modern chemistry. We will examine the key scientists who contributed to our understanding of atomic structure from Democritus to Thomson, Rutherford and Bohr. Moving on to Chapter 2 we will explore the structure of the atom in more detail including the subatomic particles that make up an atom: protons, neutrons, and electrons. We will also discuss the concept of atomic number, atomic mass, and isotopes. In Chapter 3 we will delve into the arrangement of electrons within an atom's electron cloud including energy levels, sublevels, and orbital shapes. We will also discuss the principles of electron configuration and how it impacts an atom's chemical properties. Chapter 4 will focus on the periodic table and how it organizes elements based on their atomic structure. We will explore periodic trends such as atomic radius, ionization energy, and electron affinity and discuss how these trends are related to an element's position on the table. In our final chapter we will discuss the significance of atomic structure in various chemical reactions including bonding, stoichiometry, and the behavior of gases. We will also explore the importance of understanding atomic structure in fields such as biochemistry, environmental science, and materials science. Throughout this book we will provide clear explanations, helpful illustrations, and interactive activities to reinforce key concepts. Whether you are a visual learner, a hands-on learner, or a theoretical learner, this book is designed to meet your needs. We hope that this book will not only enhance your understanding of atomic structure but also inspire you to further explore the wonders of chemistry. So let's dive in and unlock the mysteries of the atom together.

*Chemistry & Atomic Structure, Etc* John David Main SMITH, 1924    **Atomic Structure and Chemical Reactions** Nevin Katz, 2007-05    Featuring the comic strip Dr. Birdley investigates    **The Structure of the Atom** Ernest Rutherford, 1913    Elementary Atomic Structure Gordon Kemble Woodgate, 1998    Atomic Spectra and Atomic Structure Gerhard Herzberg, 1937    **Atomic Structure and Chemistry** I. Hepburn, 1926\*    Chemistry and Atomic Structure John D. Main-Smith, 1924    Atomic Structure Laying the Foundation, 2008-07    **Atomic Structure** Laying the Foundation, 2008-07    Elementary Atomic Structure G.K.

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