

Experiment 11

Data and Calculations: The Atomic Spectrum of Hydrogen

A. Calculation of the Energy Levels of the Hydrogen Atom

Excesses are to be calculated from Equations 6 for the 10 lowest energy states.

Table 11.2

B. Calculation of the Wavelengths of the Lines in the Hydrogen Atomic Spectrum

In the upper half of each box, write ΔE , the difference in energy in kJ/mole between E_{∞} and E_{eq} . In the lower half of the box, write the λ in nm associated with that value of ΔE .

Table 11.2

	6	5	4	3	2	1
1	12.75...6 KJ/mol	1250.56 KJ/mol	12.30...04 KJ/mol	11.6...25 KJ/mol	0.74...05 KJ/mol	
	93.78 nm	94.91 nm	91.15 nm	101.57 nm	121.76 nm	
2	291.57 KJ/mol	275.52 KJ/mol	240.01 KJ/mol	191.13 KJ/mol		
	410.23 nm	434.17 nm	476.26 nm	626.46 nm		
3	109.34 KJ/mol	93.5 KJ/mol	63.75 KJ/mol			
	1694.67 nm	171.03 nm	13.15 nm			
4	0.5...6 KJ/mol	29.52 KJ/mol				
	16.75...10 nm	4052.46 nm				
5	10.04 KJ/mol					
	745%...04 nm					
				$\Delta E = E_n - E_m$		
				$\lambda \text{ [nm]} = \frac{1.19627 \times 10^9}{\Delta E}$		

$$\Delta E = E_{\text{final}} - E_{\text{initial}}$$

$$\lambda \text{ [nm]} = \frac{1.19627 \times 10^9}{\Delta E}$$

(continued on following page)

Atomic Spectrum Of Hydrogen Lab Answer Key

Michael Rajnik

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