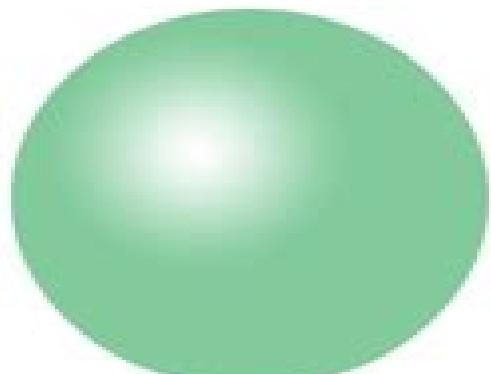


Atomic Radius

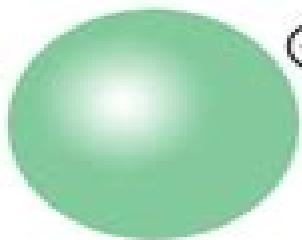
VS

Ionic Radius

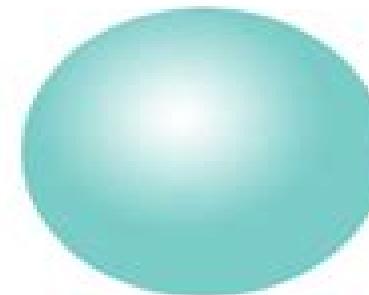
Cation: Atomic Radius > Ionic Radius



Na
190 pm



Na^+
116 pm



Mg
145 pm

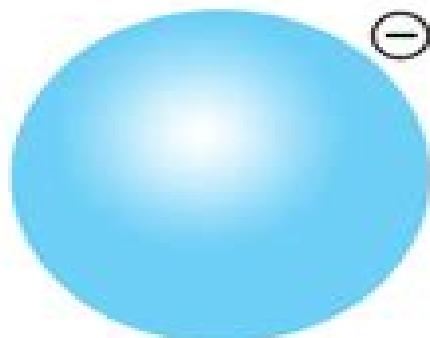


Mg^{2+}
86 pm

Anion: Atomic Radius < Ionic Radius



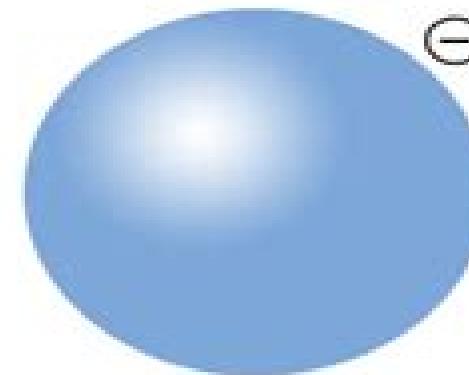
Cl
79 pm



Cl^-
167 pm



S
88 pm



S^{2-}
184 pm

Atomic And Ionic Radii Answer Key

James C. Hill

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could be processed in order to provide even more technologically effective materials for our use This increasing reliance on rocks and minerals required that more and more people be conversant with these materials and their properties The atomistic theory of the Greeks was a solely philosophical construct and further development had to await a more sophisticated approach to Science The first steps in this direction were taken by who else but Isaac Newton 1643 1727 AD Although his ideas on action at a distance initially referred to planets he also considered them as applying to atoms and concluded from physical evidence involving surface tension and viscosity that there must be strong attractions between atoms In what must be considered as insight of legendary proportions Roger Joseph Boscovich 1711 1787 a Jesuit mathematician from Croatia proposed that at very short distances atoms repulse each other the repulsion increasing indefinitely as the particles become closer together whereas at longer distances apart atoms oscillate between attraction and repulsion Frank Hawthorne uses the republication of this set of landmark papers as a vehicle to focus on the development of key issues concerning structural connectivity in inorganic solids of which minerals are a key component and to look at where we are today in our understanding of crystal structure

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