

# AQA Physics Equations Sheet

## GCSE Physics (8463)

kinetic energy = $0.5 \times \text{mass} \times (\text{speed})^2$	$E_k = \frac{1}{2} m v^2$
elastic potential energy = $0.5 \times \text{spring constant} \times (\text{extension})^2$	$E_e = \frac{1}{2} k x^2$
gravitational potential energy = mass $\times$ gravitational field strength $\times$ height	$E_p = m g h$
change in thermal energy = mass $\times$ specific heat capacity $\times$ temperature change	$\Delta E = m c \Delta \theta$
power = $\frac{\text{energy transferred}}{\text{time}}$	$P = \frac{E}{t}$
power = $\frac{\text{work done}}{\text{time}}$	$P = \frac{W}{t}$
efficiency = $\frac{\text{useful output energy transfer}}{\text{total input energy transfer}}$	
efficiency = $\frac{\text{useful power output}}{\text{total power input}}$	
charge flow = current $\times$ time	$Q = I t$
potential difference = current $\times$ resistance	$V = I R$
power = potential difference $\times$ current	$P = V I$
power = $(\text{current})^2 \times \text{resistance}$	$P = I^2 R$
energy transferred = power $\times$ time	$E = P t$
energy transferred = charge flow $\times$ potential difference	$E = Q V$
density = $\frac{\text{mass}}{\text{volume}}$	$\rho = \frac{m}{V}$

# Aqa Physics Unit Jun Paper 2005

**David Kelly**

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## **Table of Contents Aqa Physics Unit Jun Paper 2005**

1. Understanding the eBook Aqa Physics Unit Jun Paper 2005
  - The Rise of Digital Reading Aqa Physics Unit Jun Paper 2005
  - Advantages of eBooks Over Traditional Books
2. Identifying Aqa Physics Unit Jun Paper 2005
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Aqa Physics Unit Jun Paper 2005
  - User-Friendly Interface
4. Exploring eBook Recommendations from Aqa Physics Unit Jun Paper 2005
  - Personalized Recommendations
  - Aqa Physics Unit Jun Paper 2005 User Reviews and Ratings
  - Aqa Physics Unit Jun Paper 2005 and Bestseller Lists
5. Accessing Aqa Physics Unit Jun Paper 2005 Free and Paid eBooks
  - Aqa Physics Unit Jun Paper 2005 Public Domain eBooks
  - Aqa Physics Unit Jun Paper 2005 eBook Subscription Services
  - Aqa Physics Unit Jun Paper 2005 Budget-Friendly Options
6. Navigating Aqa Physics Unit Jun Paper 2005 eBook Formats

- ePub, PDF, MOBI, and More
  - Aqa Physics Unit Jun Paper 2005 Compatibility with Devices
  - Aqa Physics Unit Jun Paper 2005 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Aqa Physics Unit Jun Paper 2005
  - Highlighting and Note-Taking Aqa Physics Unit Jun Paper 2005
  - Interactive Elements Aqa Physics Unit Jun Paper 2005
8. Staying Engaged with Aqa Physics Unit Jun Paper 2005
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Aqa Physics Unit Jun Paper 2005
9. Balancing eBooks and Physical Books Aqa Physics Unit Jun Paper 2005
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Aqa Physics Unit Jun Paper 2005
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Aqa Physics Unit Jun Paper 2005
- Setting Reading Goals Aqa Physics Unit Jun Paper 2005
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Aqa Physics Unit Jun Paper 2005
- Fact-Checking eBook Content of Aqa Physics Unit Jun Paper 2005
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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 $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu(NO}_3)_2$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole H}_2\text{S} \times 2 \text{ mole H}_2\text{O} / 2 \text{ mole H}_2\text{S} = 16.6 \text{ mole H}_2\text{O}$  is your theoretical yield so percent yield  $137.1 \text{ g H}_2\text{O} / 149.4 \text{ g H}_2\text{O} \times 100 = 91.77\%$  problems 6-8 are similar to problem 5

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web a write a balanced equation for the reaction  $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0

g al x 1 mol al x 2 mol al<sub>2</sub>O<sub>3</sub> x 101 964 g al<sub>2</sub>O<sub>3</sub> 302 3 g al<sub>2</sub>O<sub>3</sub> 26 982 g al 4 mol al 1 mol al<sub>2</sub>O<sub>3</sub> c determine the percent yield yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h<sub>2</sub> are left over 13

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