

AQA Physics Equations Sheet

GCSE Physics (8463)

kinetic energy = $0.5 \times \text{mass} \times (\text{speed})^2$	$E_k = \frac{1}{2} m v^2$
elastic potential energy = $0.5 \times \text{spring constant} \times (\text{extension})^2$	$E_e = \frac{1}{2} k e^2$
gravitational potential energy = $\text{mass} \times \text{gravitational field strength} \times \text{height}$	$E_p = m g h$
change in thermal energy = $\text{mass} \times \text{specific heat capacity} \times \text{temperature change}$	$\Delta E = m c \Delta \theta$
power = $\frac{\text{energy transferred}}{\text{time}}$	$P = \frac{E}{t}$
power = $\frac{\text{work done}}{\text{time}}$	$P = \frac{W}{t}$
efficiency = $\frac{\text{useful output energy transfer}}{\text{total input energy transfer}}$	
efficiency = $\frac{\text{useful power output}}{\text{total power input}}$	
charge flow = $\text{current} \times \text{time}$	$Q = I t$
potential difference = $\text{current} \times \text{resistance}$	$V = I R$
power = $\text{potential difference} \times \text{current}$	$P = V I$
power = $(\text{current})^2 \times \text{resistance}$	$P = I^2 R$
energy transferred = $\text{power} \times \text{time}$	$E = P t$
energy transferred = $\text{charge flow} \times \text{potential difference}$	$E = Q V$
density = $\frac{\text{mass}}{\text{volume}}$	$\rho = \frac{m}{V}$

Aqa Physics Unit Jun Paper 2005

David Kelly



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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8 3 mole h₂s and multiply it with 2 mole h₂o over 2 mole h₂s from the equation

$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$

 8 3 mole h₂s x 18g h₂o i mole h₂o 149 4g h₂o is your theoretical yield so percent yield 137 1g h₂o actual yield 149 4g h₂o x 100
 91 77 problems 6 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

 if the reaction of 21 3 grams of o₂ produces a 47 3 yield how many grams of co would be produced step one figure out how many moles of the limiting reagent you have

$$\frac{21.3 \text{ grams O}_2}{32 \text{ grams mole}} = 0.66 \text{ moles O}_2$$

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$$\text{calculated mass of the product} \times 100$$

 in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction

$$\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

 4 al 3 o₂ 2 al₂o₃ b determine the theoretical yield of al₂o₃ 160 0

g al x 1 mol al x 2 mol al₂O₃ x 101 964 g al₂O₃ 302 3 g al₂O₃ 26 982 g al 4 mol al 1 mol al₂O₃ c determine the percent yield
 yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be
 weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12
 c is the limiting reagent 4 33 g of h 2 are left over 13

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 from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s
 remains when the reaction is complete

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 yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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 how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is
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 ch₃oh can be produced by the following reaction 2h₂ co ch₃oh calculate the theoretical yield of ch₃oh if 68 5 g of co is
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 which reactant is limiting assuming we started with 30 0 grams of ammonium nitrate and 50 0 grams of sodium phosphate
 what is the

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 80 g I_2O_5 1 mol I_2O_5 1 mol I_2 xs 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

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 text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3

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