

1 From AQA Chemistry Unit 1 Foundation Chemistry CHEM1 May 2014 (Question 1).

a Table 1.1 shows some data about fundamental particles in an atom.

Table 1.1 Mass of fundamental particles

Particle	proton	neutron	electron
Mass / g	$1.6725 \times 10^{-24}$	$1.6725 \times 10^{-24}$	$0.0009 \times 10^{-24}$

i An atom of hydrogen can be represented as  ${}^1\text{H}$ . Use data from Table 1.1 to calculate the mass of this hydrogen atom.

$1.6734 \times 10^{-27}$  kg

(1 mark)

ii Which one of the following is a fundamental particle that would not be deflected by an electric field?

A electron  
B neutron  
C proton

Write the correct letter, **A**, **B**, or **C**, on the line below.

**B**

(1 mark)

b A naturally occurring sample of the element boron has a relative atomic mass of 10.8. In this sample, boron exists as two isotopes,  ${}^{10}\text{B}$  and  ${}^{11}\text{B}$ .

i Calculate the percentage abundance of  ${}^{10}\text{B}$  in this naturally occurring sample of boron.

$10.8 = (10)(x) + (11)(1-x)$

$10.8 = 10x + 11 - 11x$

$10.8 - 11 = 10x - 11x$

$-0.2 = -x$

$x = 0.2 \quad 0.2 \times 100 = 20\%$

(2 marks)

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**Nora Henry, Alyn G. McFarland**

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### **Table of Contents Aqa Exam Style Questions Answers Chemistry As Chapter 15**

1. Understanding the eBook Aqa Exam Style Questions Answers Chemistry As Chapter 15
  - The Rise of Digital Reading Aqa Exam Style Questions Answers Chemistry As Chapter 15
  - Advantages of eBooks Over Traditional Books
2. Identifying Aqa Exam Style Questions Answers Chemistry As Chapter 15
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Aqa Exam Style Questions Answers Chemistry As Chapter 15
  - User-Friendly Interface
4. Exploring eBook Recommendations from Aqa Exam Style Questions Answers Chemistry As Chapter 15

- Personalized Recommendations
- Aqa Exam Style Questions Answers Chemistry As Chapter 15 User Reviews and Ratings
- Aqa Exam Style Questions Answers Chemistry As Chapter 15 and Bestseller Lists

5. Accessing Aqa Exam Style Questions Answers Chemistry As Chapter 15 Free and Paid eBooks

- Aqa Exam Style Questions Answers Chemistry As Chapter 15 Public Domain eBooks
- Aqa Exam Style Questions Answers Chemistry As Chapter 15 eBook Subscription Services
- Aqa Exam Style Questions Answers Chemistry As Chapter 15 Budget-Friendly Options

6. Navigating Aqa Exam Style Questions Answers Chemistry As Chapter 15 eBook Formats

- ePUB, PDF, MOBI, and More
- Aqa Exam Style Questions Answers Chemistry As Chapter 15 Compatibility with Devices
- Aqa Exam Style Questions Answers Chemistry As Chapter 15 Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Aqa Exam Style Questions Answers Chemistry As Chapter 15
- Highlighting and Note-Taking Aqa Exam Style Questions Answers Chemistry As Chapter 15
- Interactive Elements Aqa Exam Style Questions Answers Chemistry As Chapter 15

8. Staying Engaged with Aqa Exam Style Questions Answers Chemistry As Chapter 15

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Aqa Exam Style Questions Answers Chemistry As Chapter 15

9. Balancing eBooks and Physical Books Aqa Exam Style Questions Answers Chemistry As Chapter 15

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Aqa Exam Style Questions Answers Chemistry As Chapter 15

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Aqa Exam Style Questions Answers Chemistry As Chapter 15

- Setting Reading Goals Aqa Exam Style Questions Answers Chemistry As Chapter 15
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Aqa Exam Style Questions Answers Chemistry As Chapter 15

- Fact-Checking eBook Content of Aqa Exam Style Questions Answers Chemistry As Chapter 15
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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