

- 1 Amended from AQA Chemistry Unit 1 Foundation Chemistry CHEM1 June 2010  
(Question 1)

Fluorine forms many compounds that contain covalent bonds.

- a i State the meaning of the term *covalent bond*.

.....  
.....

(1 mark)

- ii Write an equation to show the formation of one molecule of  $\text{ClF}_3$  from chlorine and fluorine molecules.

.....  
.....

(1 mark)

- b Draw the shape of a dichlorodifluoromethane molecule ( $\text{CCl}_2\text{F}_2$ ) and the shape of a chlorine trifluoride molecule ( $\text{ClF}_3$ ). Include any lone pairs of electrons that influence the shape.

Shape of  $\text{CCl}_2\text{F}_2$

Shape of  $\text{ClF}_3$

(2 marks)

- c Suggest the strongest type of intermolecular force between  $\text{CCl}_2\text{F}_2$  molecules.

.....  
.....

(1 mark)

- d  $\text{BF}_3$  is a covalent molecule that reacts with an  $\text{F}^-$  ion to form a  $\text{BF}_4^-$  ion.

- i Name the type of bond formed when a molecule of  $\text{BF}_3$  reacts with an  $\text{F}^-$  ion. Explain how this bond is formed.

Type of bond

.....  
.....

Explanation

.....  
.....

(2 marks)

# Aqa Chemistry A2 Exam Style Question Answers

**Mark Clarkson**



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