

1 Amended from AQA Chemistry Unit 2 Chemistry in Action CHEM2 January 2011 (Question 9)

a A student carried out a laboratory experiment to determine the enthalpy change when a sample of butan-1-ol was burnt.

The student found that the temperature of 1.75 g of water increased by 8.0 °C when 5.00×10^{-3} mol of pure butan-1-ol was burnt in air and the heat produced was used to warm the water.

Use the student's results to calculate a value, in kJ mol^{-1} , for the enthalpy change when one mole of butan-1-ol is burnt.

Give your answer to an appropriate number of significant figures. (The specific heat capacity of water is $4.18 \text{ J K}^{-1} \text{ g}^{-1}$.)

f3 (maths)

Q. Give the meaning of the term standard enthalpy of combustion.

6.3 marks)

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Sue Orwin

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