

**1** Amended from AQA Chemistry Unit 2 Chemistry in Action CHEM2 January 2011  
(Question 9)

- a** A student carried out a laboratory experiment to determine the enthalpy change when a sample of butan-1-ol was burnt.  
The student found that the temperature of 175 g of water increased by 8.0 °C when  $5.00 \times 10^{-3}$  mol of pure butan-1-ol was burnt in air and the heat produced was used to warm the water.  
Use the student's results to calculate a value, in  $\text{kJ mol}^{-1}$ , for the enthalpy change when one mole of butan-1-ol is burnt.  
Give your answer to an appropriate number of significant figures.  
(The specific heat capacity of water is  $4.18 \text{ J K}^{-1} \text{ g}^{-1}$ )

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(3 marks)

- b i** Give the meaning of the term *standard enthalpy of combustion*.

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(3 marks)

# Aqa As Chemistry Exam Style Question Answers

**Sue Orwin**



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