

Exercises



- Provide an expression for K_p for reaction 1. Why does a term for C(graphite) not appear in the K_p expression for the reaction of Table 1?
- Calculate the equilibrium constant for the reaction using tabulated data.

$$\Delta_r H^\circ = -48.4 \text{ kJ/mol}$$



$$\Delta_r S^\circ = -87 \text{ J/K/mol}$$

Write the equilibrium constant for this reaction in terms of the partial pressures.

- Calculate the equilibrium constant for

$$\Delta_r G_f^\circ (\text{CaO}) = -646.3 \text{ kJ/mol}$$



$$\Delta_r G_f^\circ (\text{CO}_2) = -394.36 \text{ kJ/mol}$$

from tabulated data.

$$\Delta_r G_f^\circ (\text{CaO}) = -128.77 \text{ kJ/mol}$$

- The reaction for the formation of nitrosyl chloride



was studied at 25°C. The pressures at equilibrium were found to be

$$P_{\text{NOCl}} = 1.2 \text{ atm}$$

$$P_{\text{NO}} = 5.0 \times 10^{-2} \text{ atm}$$

$$P_{\text{Cl}_2} = 3.0 \times 10^{-1} \text{ atm}$$

- Calculate the value of K_p for this reaction at 25°C.
- If 1.0 atm of NOCl(g) was placed into an evacuated container would a reaction occur?
- If 1.0 atm each of NOCl(g), Cl₂(g) and NO(g) were placed into an evacuated container would a reaction occur? If so in which direction would the reaction proceed?
- If 1.0 atm of NOCl(g), 5.0×10^{-2} atm of NO and 3.0×10^{-1} atm of Cl₂ were placed into an evacuated container would a reaction occur? If so in which direction would the reaction proceed?

Show all calculations and explain each answer.

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