

CHEM 12 Hess's Law Worksheet

1. Calculate the ΔH for the reaction: $C_2H_4(g) + H_2(g) \rightarrow C_2H_6(g)$,
from the following data:



2. Calculate the ΔH for the reaction:
 $4 NH_3(g) + 5 O_2(g) \rightarrow 4 NO(g) + 6 H_2O(g)$, from the following data:



3. Find the ΔH_f° (heat of formation) for acetic acid, $HC_2H_3O_2$, using the following thermochemical data:



Fun Fact: C(s) is in the form of graphite.

Acetic acid, $HC_2H_3O_2$, actually has the structure: CH_3COOH



4. Calculate the ΔH for the reaction: $CH_4(g) + NH_3(g) \rightarrow HCN(g) + 3 H_2(g)$,
from the following data:



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Steven S. Zumdahl, Thomas J.. Hummel

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web we can express the first law in many ways one of the more useful expressions is that the change in internal energy ΔU of a system in any process is equal to the heat q added to the system plus the work w done on the system by its surroundings $\Delta U = q + w$

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