

CHEMISTRY MULTIPLE CHOICE QUESTIONS

A. Atoms, molecules and stoichiometry

2002 -2014

MEGA LECTURE

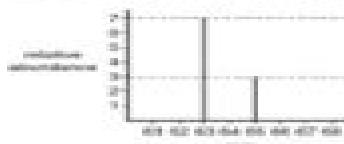
As a simplification, an adult human can be considered to have a body mass of 70,000 g of methachloro isopropyl bromide $\text{CH}_2\text{ClCH}_2\text{CH}_3\text{Br}$.

What mass of carbon dioxide does a person exhale each day if all the methachloro isopropyl bromide is oxidized and burned?

A. 0.007 kg B. 0.0004 kg C. 0.004 kg D. 0.04 kg

(2002 MCQ (2))

The diagram shows the mass spectrum of a sample of naturally-occurring copper.



What is the relative atomic mass of this copper?

A. 63.0 B. 63.5 C. 63.8 D. 64.0

(2002 MCQ (2))

Which pairs of compounds have the same empirical formula?

1. ethane and ethene
2. ethane and cyclohexane
3. cyclohexane and cyclohexene

(2002 MCQ (2))

A mixture of 10cm^3 of methane and 10cm^3 of ethane was ignited with an excess of oxygen, after cooling to room temperature, the residual gas was passed through successive palladium perls.

Which compounds have the empirical formula CH_2O ?

A. methanol
B. ethanoic acid
C. methyl ethanoate

(2002 MCQ (2))

What is the number of molecules in 500 cm^3 of oxygen under room conditions?

A. 1.25×10^{24}
B. 1.34×10^{24}
C. 3.0×10^{24}
D. 3.0×10^{25}

(2002 MCQ (2))

In the preparation of soft soap, glycerol is added.



It is usually hypothesized on that, on average, one of the hydroxyls is converted and the $\text{CH}_3\text{COCH}_2\text{CH}_2\text{OCH}_2\text{CH}_2\text{OCH}_2\text{CH}_2\text{OH}$ molecule and free hydroxyls are converted into the $\text{CH}_3\text{COCH}_2\text{CH}_2\text{OCH}_2\text{CH}_2\text{OCH}_2\text{CH}_2\text{OCH}_2\text{CH}_2\text{OH}$ ester.

How many moles of the ester are required to convert one mole of glycerol to皂素 into the soft soap?

A. 0 B. 1 C. 10 D. 100

(2002 MCQ (2))

One of the Data Blocked is relevant to this question.

Standard enthalpies of formation for selected substances are given in the table below.

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AN Whitehead

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