

# 5

## ELECTRONS IN ATOMS

### A. Matching

Match each term in Column A with the correct description in Column B. Write the letter of the correct term on the line.

| Column A                                                                                                                               | Column B                     |
|----------------------------------------------------------------------------------------------------------------------------------------|------------------------------|
| _____ 1. the ways in which electrons are arranged around the nuclei of atoms                                                           | a. Aufbau principle          |
| _____ 2. the ejection of electrons by metals when light shines on them                                                                 | b. quantum                   |
| _____ 3. the region around the nucleus of an atom where an electron is likely to be moving                                             | c. photoelectric effect      |
| _____ 4. An atomic orbital may describe at most two electrons.                                                                         | d. atomic orbitals           |
| _____ 5. the regions within which electrons have the highest probability of being found                                                | e. Pauli exclusion principle |
| _____ 6. When electrons occupy orbitals of equal energy, one electron enters each orbital until all the orbitals contain one electron. | f. energy level              |
| _____ 7. the amount of energy required to move an electron from its present energy level to the next highest one                       | g. Hund's rule               |
| _____ 8. light quanta                                                                                                                  | h. electron configurations   |
| _____ 9. Electrons enter orbitals of lowest energy first.                                                                              | i. quantum mechanical model  |
| _____ 10. the modern description of the electrons in atoms                                                                             | j. photons                   |

### B. Multiple Choice

Choose the best answer and write its letter on the line.

- \_\_\_\_\_ 11. Bohr's contribution to the development of atomic structure:
- was referred to as the "planetary model."
  - was the discovery that electrons surround a dense nucleus.
  - was proposed that electrons travel in circular orbits around the nucleus.
  - is the quantum mechanical model.
- \_\_\_\_\_ 12. What is the total number of orbitals in the third principal energy level?
- 1
  - 4
  - 9
  - 16

# 5 Electrons In Atoms Chapter 5 Test Answers

**J Ma**



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