

Name Key Class \_\_\_\_\_ Date \_\_\_\_\_

## Atomic Structure

### Subatomic Particles in Neutral Atoms

The table below contains information about several isotopes. Use the information given to fill in the blanks. Assume all atoms are neutral.

Isotope Name	Nuclear Symbol	Atomic Number	Mass Number	# of Protons	# of Neutrons	# of Electrons
1. calcium-40	${}_{20}^{40}\text{Ca}$	20	40	20	20	20
2. iron-56	${}_{26}^{56}\text{Fe}$	26	56	26	30	26
3. oxygen-18	${}_{8}^{18}\text{O}$	8	18	8	10	8
4. gold-197	${}_{79}^{197}\text{Au}$	79	197	79	118	79

### Subatomic Particles in Ions

The table below contains information about several ions. Use the information given to fill in the blanks.

Element Name	Ion Symbol	Atomic Number	Mass Number	# of Protons	# of Neutrons	# of Electrons
5. chlorine	$\text{Cl}^{-}$	17	35	17	18	18
6. silver	$\text{Ag}^{+}$	47	107	47	60	46
7. oxygen	$\text{O}^{2-}$	8	16	8	8	10
8. aluminum	$\text{Al}^{3+}$	13	27	13	14	10

### Average Atomic Mass

9. Calculate the average atomic mass for neon if its abundance in nature is 90.5% neon-20, 0.3% neon-21, and 9.2% neon-22.

$$\frac{(90.5)(20 \text{ amu}) + (0.3)(21 \text{ amu}) + (9.2)(22 \text{ amu})}{100} = 20.18 \text{ amu}$$

10. Calculate the average atomic mass of silver if 13 out of 25 atoms are silver-107 and 12 out of 25 atoms are silver-109.

$$\frac{(13)(107 \text{ amu}) + (12)(109 \text{ amu})}{25} = 107.96 \text{ amu}$$

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**Peterson's**



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