

Chem 115: Answers to Chapter 6_Aqueous Ionic Equilibrium Practice Questions

Name of Instructor: Pearson W. K. Ahiaonu

Practice Question 1

Calculate the pH of the buffer that results from mixing 60.0 mL of 0.250 M HCHO_2 and 15.0 mL of 0.500 M NaCHO_2 (K_a for $\text{HCHO}_2 = 1.8 \times 10^{-4}$)



$$\begin{aligned}\text{Total Volume of buffer} &= 60.0 \text{ mL} + 15.0 \text{ mL} \\ &= 75.0 \text{ mL} = 0.075 \text{ L}\end{aligned}$$

$$[\text{HCHO}_2] = \frac{0.060 \text{ L} \times 0.250 \text{ M}}{0.075 \text{ L}} = 0.200 \text{ M}$$

$$[\text{CHO}_2^-] = \frac{0.015 \text{ L} \times 0.500 \text{ M}}{0.075 \text{ L}} = 0.100 \text{ M}$$

$$\Rightarrow \text{pH} = \text{p}K_a + \log \frac{[\text{CHO}_2^-]}{[\text{HCHO}_2]}$$

$$= -\log (1.8 \times 10^{-4}) + \log \left(\frac{0.100}{0.200} \right)$$

$$= - (0.25527 - 4) + (-0.30103)$$

$$= 3.7447275 - 0.30103$$

$$= 3.44$$

Aqueous Equilibrium Practice Problems

**John Olmsted, Greg Williams, Robert C.
Burk**



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Solving Equilibrium Problems with Applications to Qualitative Analysis Steven S. Zumdahl, 1989

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Carbonate Chemistry of Aquatic Systems Richard E. Loewenthal, Gerrit van Rooyen Marais, 1976 **Student's Guide, Chemistry, the Central Science** James C. Hill, 1991

Problem Exercises for General Chemistry G. Gilbert

Long, Forrest C. Hentz, 1986-08-14 A question answer supplement to accompany a 1 year general chemistry course This text

is a multiple choice format and can be used with any standard general chemistry text The exercises emphasize the importance of problem solving and equation writing in the style used in general chemistry examinations and professional school aptitude examinations such as the MCAT and VCAT Chemistry John W. Moore, Conrad L. Stanitski, Peter C. Jurs, 2002 CHEMISTRY THE MOLECULAR SCIENCE is intended to help students develop a broad overview of chemistry and chemical reactions an understanding of the most important concepts and models that chemists and those in chemistry related fields use an appreciation of the many ways chemistry impacts our daily lives the ability to apply the facts concepts and models of chemistry appropriately to new situations in chemistry other sciences and engineering and to other disciplines

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Principles, Fourth Edition, by Dickerson, Gray, Darensbourg, and Darensbourg Patricia L. Samuel, 1984 *Acta chimica*, 1971 Nr 64 ladkowska J Polyn mes quasi univalents et univalents 1960 **Principles of Chemistry** Loren G. Hepler, Wayne L. Smith, 1975 Proceedings of the World Geothermal Congress, 1995, Florence, Italy, 18-31 May 1995, 1995 Four hundred and fifty four papers arranged in 15 sections **Chemical principles** Paul B. Kelter, 1998

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