

Chem 115: Answers to Chapter 6_Aqueous Ionic Equilibrium Practice Questions

Name of Instructor: Pearson W. K. Ahiahouu

Practice Question 1

Calculate the pH of the buffer that results from mixing 60.0 mL of 0.250 M HCHO_2 and 15.0 mL of 0.500 M NaCHO_2 (K_a for $\text{HCHO}_2 = 1.8 \times 10^{-4}$)



$$\begin{aligned} \text{Total Volume of buffer} &= 60.0 \text{ mL} + 15.0 \text{ mL} \\ &= 75.0 \text{ mL} = 0.075 \text{ L} \end{aligned}$$

$$[\text{HCHO}_2] = \frac{0.060 \text{ L}}{0.075 \text{ L}} \times 0.250 \text{ M} = 0.200 \text{ M}$$

$$[\text{CHO}_2^-] = \frac{0.015 \text{ L}}{0.075 \text{ L}} \times 0.500 \text{ M} = 0.100 \text{ M}$$

$$\begin{aligned} \Rightarrow \text{pH} &= \text{p}K_a + \log \frac{[\text{CHO}_2^-]}{[\text{HCHO}_2]} \\ &= -\log (1.8 \times 10^{-4}) + \log \left(\frac{0.100}{0.200} \right) \\ &= -(0.25527 - 4) + (-0.30103) \\ &= 3.7447275 - 0.30103 \\ &= 3.44 \end{aligned}$$

Aqueous Equilibrium Practice Problems

John Olmsted, Greg Williams, Robert C. Burk

Aqueous Equilibrium Practice Problems:

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Chemistry Steven S. Zumdahl, 1997 **CHEMISTRY** SILBERBERG, 2003

Carbonate Chemistry of Aquatic Systems Richard E. Loewenthal, Gerrit van Rooyen Marais, 1976 **Student's Guide, Chemistry, the Central Science** James C. Hill, 1991 **Problem Exercises for General Chemistry** G. Gilbert Long, Forrest C. Hentz, 1986-08-14 A question answer supplement to accompany a 1 year general chemistry course. This text

is a multiple choice format and can be used with any standard general chemistry text The exercises emphasize the importance of problem solving and equation writing in the style used in general chemistry examinations and professional school aptitude examinations such as the MCAT and VCAT Chemistry John W. Moore,Conrad L. Stanitski,Peter C. Jurs,2002 CHEMISTRY THE MOLECULAR SCIENCE is intended to help students develop a broad overview of chemistry and chemical reactions an understanding of the most important concepts and models that chemists and those in chemistry related fields use an appreciation of the many ways chemistry impacts our daily lives the ability to apply the facts concepts and models of chemistry appropriately to new situations in chemistry other sciences and engineering and to other disciplines

Chemical Principles in Practice Jerry A. Bell,1967 This forward looking laboratory text includes thirty three experiments designed to serve high level introductory chemistry courses It is intended in part to give the students practice in techniques not only those generally considered the techniques of chemistry but also those of mathematics However the only prerequisite is a good background in high school mathematics Ideally suited for use with University Chemistry or College Chemistry by Bruce Mahan this book follows the physical chemistry approach of these and most other modern chemistry texts maintaining the same flavor emphasis and degree of sophistication All experiments are open ended and include suggestions for following them up There is considerable emphasis on the involvement of the student in planning his own procedures and quantitative thinking Publisher General Chemistry James P. Birk,Donald M. Baird,1994-03 **Study Guide for Chemical Principles, Fourth Edition, by Dickerson, Gray, Darensbourg, and Darensbourg** Patricia L. Samuel,1984 *Acta chimica* ,1971 Nr 64 laskowska J Polyn mes quasi univalents et univalents 1960 **Principles of Chemistry** Loren G. Hepler,Wayne L. Smith,1975 Proceedings of the World Geothermal Congress, 1995, Florence, Italy, 18-31 May 1995 ,1995 Four hundred and fifty four papers arranged in 15 sections **Chemical principles** Paul B. Kelter,1998

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Table of Contents Aqueous Equilibrium Practice Problems

1. Understanding the eBook Aqueous Equilibrium Practice Problems
 - The Rise of Digital Reading Aqueous Equilibrium Practice Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Aqueous Equilibrium Practice Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Aqueous Equilibrium Practice Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Aqueous Equilibrium Practice Problems
 - Personalized Recommendations

- Aqueous Equilibrium Practice Problems User Reviews and Ratings
- Aqueous Equilibrium Practice Problems and Bestseller Lists

5. Accessing Aqueous Equilibrium Practice Problems Free and Paid eBooks

- Aqueous Equilibrium Practice Problems Public Domain eBooks
- Aqueous Equilibrium Practice Problems eBook Subscription Services
- Aqueous Equilibrium Practice Problems Budget-Friendly Options

6. Navigating Aqueous Equilibrium Practice Problems eBook Formats

- ePUB, PDF, MOBI, and More
- Aqueous Equilibrium Practice Problems Compatibility with Devices
- Aqueous Equilibrium Practice Problems Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Aqueous Equilibrium Practice Problems
- Highlighting and Note-Taking Aqueous Equilibrium Practice Problems
- Interactive Elements Aqueous Equilibrium Practice Problems

8. Staying Engaged with Aqueous Equilibrium Practice Problems

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Aqueous Equilibrium Practice Problems

9. Balancing eBooks and Physical Books Aqueous Equilibrium Practice Problems

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Aqueous Equilibrium Practice Problems

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Aqueous Equilibrium Practice Problems

- Setting Reading Goals Aqueous Equilibrium Practice Problems
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Aqueous Equilibrium Practice Problems

- Fact-Checking eBook Content of Aqueous Equilibrium Practice Problems

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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