

Naming Ionic Compounds

What are the structural units that make up ionic compounds and how are they named?

Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl . In talking about chemistry however, it is a bit tacky to say "em-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

Model 1 – Ion Charges for Selected Elements

1	H ⁺														
2	Li ⁺	Be ²⁺									N ³⁻	O ²⁻	F ⁻		
3	Na ⁺	Mg ²⁺	Transition elements						Al ³⁺		P ³⁻	S ²⁻	Cl ⁻		
4	K ⁺	Ca ²⁺	Fe ²⁺	Ni ²⁺	Cu ⁺	Zn ²⁺							Br ⁻		
5	Rb ⁺	Sr ²⁺			Ag ¹⁺			Sn ²⁺					I ⁻		
6		Ba ²⁺				Hg ₂ ²⁺		Pb ²⁺							
	Cations										Anions				

4. 1. Based on the information in Model 1:
- Identify three elements that form only one cation.
Li, Na, K, Rb, Be, Mg, Ca, Sr, Ba, Ag, Zn, Al (H according to this sheet, but technically not)
 - Identify three elements that form only one anion.
N, P, O, S, F, Cl, Br, I
 - Identify three elements that form more than one cation.
Fe, Ni, Cu, Hg, Sn, Pb
 - In what region of the periodic table are these "multiple ion" elements usually located?
Center, transition elements (metals)
1. 2. Consider the ions of potassium (K) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of potassium (K) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.
 K_2S
2. 3. Consider the ions of iron (Fe) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.



Answer Key Ionic Bonds Ionic Compounds

Ann Daniels

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 - The Rise of Digital Reading Answer Key Ionic Bonds Ionic Compounds
 - Advantages of eBooks Over Traditional Books
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 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
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 - User-Friendly Interface
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 - Answer Key Ionic Bonds Ionic Compounds Enhanced eBook Features
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 - Interactive Elements Answer Key Ionic Bonds Ionic Compounds
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Ionic Bonds Ionic Compounds
- 9. Balancing eBooks and Physical Books Answer Key Ionic Bonds Ionic Compounds
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Ionic Bonds Ionic Compounds
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key Ionic Bonds Ionic Compounds
 - Setting Reading Goals Answer Key Ionic Bonds Ionic Compounds
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key Ionic Bonds Ionic Compounds
 - Fact-Checking eBook Content of Answer Key Ionic Bonds Ionic Compounds
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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