

## Mole Conversions Worksheet #1

1. Mole  $\rightarrow$  Mass Conversions – using molar mass of each substance, convert the following quantities.

a. 10.0 mol Cr <b>520 g</b>	f. 0.160 mol H <sub>2</sub> O <b>2.88 g</b>
b. 3.32 mol K <b>130 g</b>	g. 5.08 mol Ca(NO <sub>3</sub> ) <sub>2</sub> <b>834 g</b>
c. $2.20 \times 10^{-3}$ mol Sn <b>0.261 g</b>	h. 15.0 mol H <sub>2</sub> SO <sub>4</sub> <b>1470 g</b>
d. 0.720 mol Be <b>6.48 g</b>	i. $4.52 \times 10^{-2}$ mol C <sub>2</sub> H <sub>6</sub> <b>1.27 \times 10^{-3} g</b>
e. 2.40 mol N <sub>2</sub> <b>67.2 g</b>	j. 0.0112 mol K <sub>2</sub> CO <sub>3</sub> <b>1.55 g</b>

2. Mass  $\rightarrow$  Mole Conversions – using molar mass of each substance convert the following quantities.

a. 72.0 g Ar <b>1.80 mol</b>	f. 27.4 g NO <sub>2</sub> <b>0.996 mol</b>
b. $3.70 \times 10^{-3}$ g B <b>3.43 \times 10^{-2} mol</b>	g. 5.00 g H <sub>2</sub> <b>2.50 mol</b>
c. 187 g Al <b>6.93 mol</b>	h. $2.64 \times 10^{-4}$ g Li <sub>2</sub> PO <sub>4</sub> <b>2.28 \times 10^{-6} mol</b>
d. 3.33 g SnF <sub>2</sub> <b>2.13 mol</b>	i. 11.0 g CH <sub>4</sub> <b>0.688 mol</b>
e. $7.21 \times 10^{-2}$ g He <b>1.80 \times 10^{-2} mol</b>	j. 847 g (NH <sub>4</sub> ) <sub>2</sub> CO <sub>3</sub> <b>8.82 mol</b>

3. What is the volume of the following gases?

a. 5.40 mol O <sub>2</sub> <b>121 L</b>
b. $3.20 \times 10^{-2}$ mol CO <sub>2</sub> <b>0.717 L</b>
c. 0.960 mol SO <sub>2</sub> <b>21.5 L</b>

4. How many moles are in each of the following volumes?

a. 89.6 L Ne <b>4.00 mol</b>
b. $1.00 \times 10^3$ L C <sub>2</sub> H <sub>6</sub> <b>44.6 mol</b>
c. $5.42 \times 10^4$ F <sub>2</sub> <b>2.42 \times 10^5 mol</b>

5. Find the number of moles in each of the number of representative particles.

a. $1.20 \times 10^{23}$ atoms of P <b>19.9 mol</b>
b. $3.87 \times 10^{23}$ molecules of AlF <sub>3</sub> <b>6.43 \times 10^{-2} mol</b>
c. $4.81 \times 10^{24}$ molecules of NH <sub>3</sub> <b>7.99 \times 10^{10} mol</b>

6. How many representative particles are in each of the following mole quantities?

a. 1.24 mol Cl <sub>2</sub> <b>7.46 \times 10^{23} molecules</b>
b. $4.20 \times 10^{-2}$ mol K <sub>2</sub> S <b>2.53 \times 10^{22} molecules</b>
c. 34.02 mol Ca(OH) <sub>2</sub> <b>2.048 \times 10^{25} molecules</b>

7. Convert the following two-step quantities, converting first to moles and then to the desired quantity.

a. Find the number of molecules in 60.0 g of N <sub>2</sub> O. <b><math>8.21 \times 10^{23}</math> molecules</b>
b. Find the volume of $3.24 \times 10^{22}$ molecules of Ne. <b>1.21 L</b>
c. Find the mass of 18.0 L of CH <sub>4</sub> . <b>12.9 g</b>
d. Find the volume of 835 g of SO <sub>2</sub> . <b>234 L</b>
e. Find the mass of one atom of nickel. <b><math>1 \times 10^{-27}</math> g</b>

# Answer Key Mass Mole Conversion Worksheet

**Julie C. Gilbert**

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