

# ANSWER KEY - ORGANIZED BY REACTION TYPE

| Practice Bb                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                 | Practice Bbb                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                         | Practice Balancing Chemical Equations                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                |
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| <b>Part A</b><br>Instructions: Balance each of the following chemical equations. Write the correct coefficients in front of the number of atoms for each element. Do not change subscripts.<br><ol style="list-style-type: none"> <li><math>1. \text{C} + \text{O}_2 \rightarrow \text{CO}_2</math><br/>Balanced Equation</li> <li><math>1. \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}</math><br/>Balanced Equation</li> <li><math>1. \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}</math><br/>Balanced Equation</li> <li><math>1. \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}</math><br/>Balanced Equation</li> <li><math>1. \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}</math><br/>Balanced Equation</li> </ol> | <b>Part B</b><br>Instructions: Balance each of the following chemical equations. Write the correct coefficients in front of the number of atoms for each element. Do not change subscripts.<br><ol style="list-style-type: none"> <li><math>1. \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}</math><br/>Balanced Equation</li> <li><math>1. \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}</math><br/>Balanced Equation</li> <li><math>1. \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}</math><br/>Balanced Equation</li> <li><math>1. \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}</math><br/>Balanced Equation</li> <li><math>1. \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}</math><br/>Balanced Equation</li> </ol> | <b>Part C: Double Replacement Reactions</b><br>Instructions: Balance each of the following chemical equations. Write the correct coefficients in front of the number of atoms for each element. Do not change subscripts.<br><ol style="list-style-type: none"> <li><math>1. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math><br/>Balanced Equation</li> <li><math>1. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math><br/>Balanced Equation</li> <li><math>1. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math><br/>Balanced Equation</li> <li><math>1. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math><br/>Balanced Equation</li> <li><math>1. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math><br/>Balanced Equation</li> </ol> |

| Practice Bb                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                            | Practice Balancing Chemical Equations                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                |
|----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| <b>Part C: Single Replacement Reactions</b><br>Instructions: Balance each of the following chemical equations. Write the correct coefficients in front of the number of atoms for each element. Do not change subscripts.<br><ol style="list-style-type: none"> <li><math>1. \text{Zn} + \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2</math><br/>Balanced Equation</li> <li><math>1. \text{Zn} + \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2</math><br/>Balanced Equation</li> <li><math>1. \text{Zn} + \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2</math><br/>Balanced Equation</li> <li><math>1. \text{Zn} + \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2</math><br/>Balanced Equation</li> <li><math>1. \text{Zn} + \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2</math><br/>Balanced Equation</li> </ol> | <b>Part D: Double Replacement Reactions</b><br>Instructions: Balance each of the following chemical equations. Write the correct coefficients in front of the number of atoms for each element. Do not change subscripts.<br><ol style="list-style-type: none"> <li><math>1. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math><br/>Balanced Equation</li> <li><math>1. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math><br/>Balanced Equation</li> <li><math>1. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math><br/>Balanced Equation</li> <li><math>1. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math><br/>Balanced Equation</li> <li><math>1. \text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}</math><br/>Balanced Equation</li> </ol> |

## ANSWER KEY

### Part A: Combustion Reactions

- $1. \text{C} + \text{O}_2 \rightarrow \text{CO}_2$
- $1. \text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- $1. \text{C}_2\text{H}_6 + 7\text{O}_2 \rightarrow 2\text{CO}_2 + 6\text{H}_2\text{O}$
- $1. \text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$
- $1. \text{C}_4\text{H}_{10} + 13\text{O}_2 \rightarrow 4\text{CO}_2 + 10\text{H}_2\text{O}$
- $1. \text{C}_5\text{H}_{12} + 8\text{O}_2 \rightarrow 5\text{CO}_2 + 6\text{H}_2\text{O}$
- $1. \text{C}_6\text{H}_{14} + 9\text{O}_2 \rightarrow 6\text{CO}_2 + 7\text{H}_2\text{O}$
- $1. \text{C}_7\text{H}_{16} + 11\text{O}_2 \rightarrow 7\text{CO}_2 + 8\text{H}_2\text{O}$
- $1. \text{C}_8\text{H}_{18} + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 9\text{H}_2\text{O}$
- $1. \text{C}_9\text{H}_{20} + 14\text{O}_2 \rightarrow 9\text{CO}_2 + 10\text{H}_2\text{O}$

### Part B: Synthesis Reactions

- $1. \text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
- $1. \text{H}_2 + \text{S}_8 \rightarrow 2\text{H}_2\text{S}$
- $1. 2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- $1. 4\text{H} + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- $1. 2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- $1. \text{Fe} + \text{S} \rightarrow \text{FeS}$
- $1. 2\text{Ca} + \text{O}_2 \rightarrow 2\text{CaO}$
- $1. 2\text{Zn} + \text{S} \rightarrow 2\text{ZnS}$
- $1. 2\text{H} + \text{Br}_2 \rightarrow 2\text{HBr}$
- $1. 4\text{H} + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

### Part C: Decomposition Reactions

- $1. 2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$
- $1. \text{H}_2\text{CO}_3 \rightarrow \text{H}_2\text{O} + \text{CO}_2$
- $1. 2\text{HClO}_2 \rightarrow 2\text{HCl} + \text{O}_2$
- $1. \text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
- $1. 2\text{HCO}_2\text{H} \rightarrow 2\text{HCl} + \text{O}_2$
- $1. \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{O} + \text{SO}_3$
- $1. \text{H}_2\text{CO}_3 \rightarrow \text{H}_2\text{O} + \text{CO}_2$
- $1. 2\text{HClO}_2 \rightarrow 2\text{HCl} + \text{O}_2$
- $1. 2\text{HClO}_2 \rightarrow 2\text{HCl} + \text{O}_2$
- $1. 2\text{HClO}_2 \rightarrow 2\text{HCl} + \text{O}_2$

### Part B: Single Replacement Reactions

- $1. \text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
- $1. \text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$
- $1. \text{Cl}_2 + 2\text{KI} \rightarrow 2\text{KCl} + \text{I}_2$
- $1. \text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
- $1. \text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
- $1. 2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{H}_2$
- $1. \text{Br}_2 + 2\text{NaI} \rightarrow 2\text{NaBr} + \text{I}_2$
- $1. \text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$
- $1. \text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$
- $1. \text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$

### Part B: Double Replacement Reactions

- $1. \text{Na}_2\text{CO}_3 + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{H}_2\text{CO}_3$
- $1. \text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$
- $1. \text{BaCl}_2 + \text{Na}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + 2\text{NaCl}$
- $1. \text{Pb(NO}_3)_2 + 2\text{KI} \rightarrow \text{PbI}_2 + 2\text{KNO}_3$
- $1. \text{CaCl}_2 + \text{Na}_2\text{CO}_3 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$
- $1. 2\text{HCl} + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{SO}_4 + 2\text{HCl}$
- $1. \text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- $1. \text{FeCl}_3 + 3\text{NaOH} \rightarrow \text{Fe(OH)}_3 + 3\text{NaCl}$
- $1. 2\text{HNO}_3 + \text{BaCl}_2 \rightarrow 2\text{HCl} + \text{Ba(NO}_3)_2$
- $1. \text{CuSO}_4 + 2\text{NaOH} \rightarrow \text{Cu(OH)}_2 + \text{Na}_2\text{SO}_4$

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# Answer Key For Grade 7 Ready Practice

**Zhe Xu**



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