

Answer Key

Chapter 10

Practice Problems

- 9.45×10^{24} molecules CO_2
- 3.27×10^{22} molecules H_2O
- 1.53 mol Fe
- 0.0341 mol sucrose
- 1.08×10^{21} formula units CuSO_4
- 839 g Sb
- 5.53 g Se
- 6.95 mol S
- 12.5 mol He
- 0.0213 mol Ni
- 6.17×10^{21} atoms Pt
- 1.88×10^{28} atoms S
- 39.7 g Hg
- 6.35×10^{-3} g
- 4.36×10^{16} Pb ions
- 0.220 mol C_6H_6
- 2.856 mol KClO_3
- 40.2 g SnSO_4
- 3.40×10^2 g (340 g expressed to 3 significant figures)
- 13.1 g Zn and 37.5 g $\text{Cu}(\text{NO}_3)_2$
- 52.93% Al, 47.07% O
- 16.39% Mg, 18.89% N, 64.72% O
- 39.17% O
- 49.30% N
- The percent compositions will be the same because the ratio of carbon atoms to hydrogen atoms is the same, 1:1.
- CH_2O
- ZnN_2O_6
- Al_4C_3
- $\text{C}_3\text{H}_4\text{O}_3$
- $\text{C}_8\text{H}_8\text{N}_2\text{O}_2$
- $\text{B}_3\text{H}_6\text{N}_3$
- $\text{Ag}_2\text{C}_2\text{O}_4$
- P_2S_5 ; same as empirical formula
- $\text{C}_9\text{H}_{12}\text{N}_6$
- $\text{CeI}_3 \cdot 9\text{H}_2\text{O}$
- $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$
- $\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_2 \cdot 3\text{H}_2\text{O}$
- $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$

Answer Key For Chemistry Ch 16 Assessment

Richard H. Langley, John Moore



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