

ACIDS AND EQUILIBRIUM

- Equilibrium constant can be used for the ionization reaction to express the extent to which a weak acids ionizes.



$$K_a = \frac{[H_3O^+][A^-]}{[HA]}$$

or

$$K_a = \frac{[H^+][A^-]}{[HA]}$$

- A smaller K_a value meaning the degree of the dissociation of the acid is low



17 Acid Base Equilibrium Packet

L Towne



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