

**16.4****CALCULATIONS INVOLVING  
COLLIGATIVE PROPERTIES****Section Review****Objectives**

- Calculate the molality and mole fraction of a solution.
- Describe how the freezing-point depression and boiling-point elevation are related to molality.

**Vocabulary**

- molality ( $m$ )
- mole fraction
- molal freezing-point depression constant ( $K_f$ )
- molal boiling-point elevation constant ( $K_b$ )

**Key Equations**

- Molality =  $\frac{\text{moles of solute}}{\text{kilogram of solvent}}$
- mole fractions:  $X_A = \frac{n_A}{n_A + n_B}$        $X_B = \frac{n_B}{n_A + n_B}$   
 where  $n_A$  = moles of solute  
 $n_B$  = moles of solvent
- $\Delta T_b = K_b \times m$
- $\Delta T_f = K_f \times m$

**Part A Completion**

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Molality is an expression of concentration involving the ratio of \_\_\_\_\_ 1 \_\_\_\_\_ particles to \_\_\_\_\_ 2 \_\_\_\_\_ particles. Molality is expressed as moles of solute per \_\_\_\_\_ 3 \_\_\_\_\_ of solvent.

Another expression of concentration is \_\_\_\_\_ 4 \_\_\_\_\_, in which concentrations are expressed as the ratio of moles of solute to the total number of moles of solvent and solute. Each solvent has a characteristic \_\_\_\_\_ 5 \_\_\_\_\_ elevation constant and molal freezing-point \_\_\_\_\_ 6 \_\_\_\_\_ constant. The elevation in boiling point of a solution can be calculated by multiplying the \_\_\_\_\_ 7 \_\_\_\_\_ concentration of the solution by the boiling-point \_\_\_\_\_ 8 \_\_\_\_\_ constant of the solvent.

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# 16 4 Calculations Involving Colligative Properties

## Answer Sheet

**Oswaal Editorial Board**



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