

Problem Set 7: Acids and Bases Answers

10. Problem Level 2

a.

Write down conditions for HNO_3 and H_2SO_4 in the solution in water: $\text{HNO}_3 \rightarrow \text{H}^+ + \text{NO}_3^-$, corresponding to $\text{pH} = 1.0$.

b.

The equilibrium is:



For HNO_2 , $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ and for H_2O , $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$, $K_w = 1.0 \times 10^{-14}$.



The charge conservation of the water and nitrogen from the equilibrium $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ and $\text{HNO}_2 \rightleftharpoons \text{H}^+ + \text{NO}_2^-$ can be written as:

$$[\text{H}^+] = [\text{NO}_2^-] + [\text{OH}^-] \quad \text{and} \quad [\text{H}^+] = [\text{NO}_2^-]$$

and the mass balance is given by

$$[\text{H}^+] + [\text{HNO}_2] + [\text{NO}_2^-] = 0.050$$

Insert the equilibrium conditions into the mass balance and simplify using the condition $[\text{H}^+] = [\text{NO}_2^-]$. The equilibrium concentration of the nitrogen species can be found:

$$[\text{HNO}_2] + 2[\text{H}^+] = 0.050 \quad \text{and} \quad [\text{HNO}_2] + 2[\text{H}^+] = 0.050$$

Divide $[\text{H}^+] = [\text{NO}_2^-]$ by the mass balance $[\text{H}^+] + [\text{HNO}_2] + [\text{NO}_2^-] = 0.050$ and simplify the charge balance $[\text{H}^+] = [\text{NO}_2^-]$ to find the concentration of the nitrogen species and the equilibrium concentration:

$$\frac{[\text{H}^+]^2}{[\text{HNO}_2]} = 4.5 \times 10^{-4}$$

Insert $[\text{H}^+] = [\text{NO}_2^-]$ into $[\text{H}^+] + [\text{HNO}_2] + [\text{NO}_2^-] = 0.050$ and solve for $[\text{H}^+] = 0.010$.

Answers To Problem Set 2

**James N. Thompson, Jr, Jenna J.
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