

# Thermochemistry Worksheet Answers

## Useful Values:

Heat of Fusion for Water:  $\Delta H_{\text{fus}} = 6 \text{ kJ/mol}$

Common error: not converting to moles

Heat of Vaporization for Water:  $\Delta H_{\text{vap}} = 45 \text{ kJ/mol}$

Common error: same

1 calorie = 4.18 Joules

Material	Specific Heat in Joules/g/K	in Cal/g/K	
Water (ice)	2.05	0.490	
Water (liquid)	4.18	1.00	Common error here: using moles instead of grams.
Water (gas)	2.08	0.498	

1. How much energy does it take to warm the following substances the listed amount?

a. 20g of ice from  $-45^\circ \text{ C}$  to  $-15^\circ \text{ C}$ ? **294 calories / 1230 Joules**  
b. 20g of water from  $25^\circ \text{ C}$  to  $55^\circ \text{ C}$ ? **600 calories / 2508 Joules**  
c. 20g of water vapor from  $130^\circ \text{ C}$  to  $160^\circ \text{ C}$ ? **299 calories / 1249 Joules**

2. To melt ice, one must bring ice to the freezing point ( $0^\circ \text{ C}$ ), and then thaw it into water. How much heat would it take to melt...

a. 30g of ice at  $0^\circ \text{ C}$ ? **10 kJ** (**10,000 J**)  
b. 30g of ice at  $-40^\circ \text{ C}$ ? **12.46 kJ** (**12,460 J**)

3. To boil water, one must raise it to the boiling point ( $100^\circ \text{ C}$ ), and then boil it into steam. How much heat would it take to boil...

a. 40g of water at  $100^\circ \text{ C}$ ? **100 kJ** (**100,000 Joules**)  
b. 40g of water at  $90^\circ \text{ C}$ ? **101.7 kJ** (**101,672 Joules**)  
c. 40g of water at  $0^\circ \text{ C}$ ? **116.7 kJ** (**116,720 Joules**)  
d. 40g of ice at  $0^\circ \text{ C}$ ? **130.1 kJ** (**130,053 Joules**)  
e. 40g of ice at  $-80^\circ \text{ C}$ ? **136.6 kJ** (**136,613 Joules**)

4. If 700 calories of heat raise 400g of Phonium by  $20^\circ \text{ C}$ , what is the specific heat of Phonium? **0.0075 cal/g/K**

5. How much heat energy is stored in 2g of ice at  $-30^\circ \text{ C}$ ? (remember: use Kelvin) **2.38 cal / 996 J**

6. How much heat energy is stored in 2g of Phonium at  $70^\circ \text{ C}$ ? (Assume Phonium is a solid at  $70^\circ \text{ C}$ ) **1.672 cal**

7. If 100g of water at  $10^\circ \text{ C}$  comes in contact with \_\_\_\_\_, what will the final temperature be?  
(Remember, you have the specific heat of Phonium from question 4.)

a. 100g of water at  $90^\circ \text{ C}$  **50° C**  
b. 100g of Phonium at  $90^\circ \text{ C}$  (Hint: the water and Phonium end at the same temperature) **16.44° C**  
Common error: 6.44 degrees is the temperature change, and the water gets warmer.  
c. 20g of Phonium at  $90^\circ \text{ C}$  **11.38° C**  
d. 1g of Phonium at  $90^\circ \text{ C}$  **10.87° C**

8. If 450g of Phakelite at  $100^\circ \text{ C}$  is added to 1000g of water at  $20^\circ \text{ C}$ , the temperature of the water increases by  $5^\circ$ . What is the specific heat of Phakelite? **0.148 cal/g/K**

9. If 450g of Nonsenseum at  $500^\circ \text{ C}$  is added to 1000g of water at  $20^\circ \text{ C}$ , the temperature of the water increases by  $25^\circ$ . What is the specific heat of Nonsenseum? **0.122 cal/g/K**

# Answers Thermochemistry Problems Worksheet Number Two

**Melonie Anne Teichert**

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