

11 Which compound has the greatest percent composition by mass of sulfur?

way
#41
(1) BaS
(2) CaS
(3) MgS
(4) SrS

way
#42
 $2 \text{HBr} \rightarrow \text{H}_2 + \text{Br}_2 \text{ (RP)}$ ←
Two molecules of HBr collide and then form H₂ and Br₂. During the collision, the bonds in the HBr molecules are

way
#43
(1) broken as energy is absorbed
(2) broken as energy is released
(3) formed as energy is absorbed
(4) formed as energy is released

BAKE

way
#44
13 Which atom in the ground state has a stable electron configuration?

way
#45
(1) carbon
(2) magnesium
(3) neon
(4) oxygen

way
#46
14 Which statement describes a multiple covalent bond?

way
#47
(1) Two electrons are shared.
(2) Four electrons are shared.
(3) Two electrons are transferred.
(4) Four electrons are transferred.

way
#48
15 The electronegativity difference between the atoms in a molecule of HCl can be used to determine

way
#49
(1) the entropy of the atoms
(2) the atomic number of the atoms
(3) the first ionization energy of the atoms
(4) the polarity of the bond between the two atoms

way
#50
16 Which two gases can not be broken down by chemical means?

way
#51
(1) CO and He
(2) CO and NH₃
(3) Xe and He
(4) Xe and NH₃

elements
Table
S

way
#52
17 Two substances in a mixture differ in density and particle size. These properties can be used to

way
#53
(1) separate the substances
(2) chemically combine the substances
(3) determine the freezing point of the mixture
(4) predict the electrical conductivity of the mixture

18 Which unit is used to express an amount of thermal energy?

way
#54
(1) gram
(2) mole
(3) joule
(4) pascal

Table
D

19 Under which conditions of temperature and pressure does a real gas behave most like an ideal gas?

way
#55
(1) low temperature and low pressure
(2) low temperature and high pressure
(3) high temperature and low pressure
(4) high temperature and high pressure

20 According to the kinetic molecular theory for an ideal gas, all gas particles

way
#56
(1) are in random, constant, straight-line motion
(2) are separated by very small distances relative to their sizes
(3) have strong intermolecular forces
(4) have collisions that decrease the total energy of the system

21 Which mathematical expression represents the heat of reaction for a chemical reaction?

way
#57
(1) (the heat of fusion) – (the heat of vaporization)
(2) (the heat of vaporization) – (the heat of fusion)
(3) (the potential energy of the products) – (the potential energy of the reactants)
(4) (the potential energy of the reactants) – (the potential energy of the products)

22 At 101.3 kPa and 298 K, a 1.0-mole sample of which compound absorbs the greatest amount of heat as the entire sample dissolves in water?

way
#58
(1) LiBr – 47.93
(2) NaOH – 44.51
(3) NaCl + 3.78
(4) NH₄Cl + 41.78

Table
I

23 For a reaction at equilibrium, which change can increase the rates of the forward and reverse reactions?

(1) a decrease in the concentration of the reactants
(2) a decrease in the surface area of the products
(3) an increase in the temperature of the system
(4) an increase in the activation energy of the forward reaction

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