

Empirical and Molecular Formulas Worksheet

Key

Objectives:

- be able to calculate empirical and molecular formulas

Empirical Formula

1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.521g of Oxygen?

$$\text{C} \quad \frac{0.783 \text{ g C}}{12.01 \text{ g/mol}} = \frac{0.065}{1.01} \sim 1.96 \quad 2$$

$$\text{H} \quad \frac{0.196 \text{ g H}}{1.01 \text{ g/mol}} = \frac{0.194}{1.01} \sim 1.96 \quad 6$$

$$\text{O} \quad \frac{0.521 \text{ g O}}{16.00 \text{ g/mol}} = \frac{0.032}{1.01} \sim 1.96 \quad 1$$



2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?

$$\text{Au} \quad \frac{89.14\%}{196.97} = \frac{0.45}{0.45} = 1 \quad \text{AuO}_{1.5}$$

$$\text{O} \quad \frac{10.80\%}{16.00} = \frac{0.675}{0.45} = 1.5$$



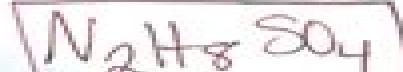
3) What is empirical formula if compound consists of 21.2%N, 6.19%H, 24.2%S and 48.5%O?

$$\text{N} \quad \frac{21.2\%}{14.01} = 1.5 \quad 2$$

$$\text{H} \quad \frac{6.19\%}{1.01} = 6.03 \quad 8$$

$$\text{S} \quad \frac{24.2\%}{32.06} = 0.75 \quad 1$$

$$\text{O} \quad \frac{48.5\%}{16.00} = 3.03 \quad 4$$



Molecular Formula

4) Empirical formula of a substance is CH_2O . Molar mass is 180. What is the molecular formula?

$$\text{EFM} = (12 + 2 \cdot 16 + 16) =$$

$$\frac{180}{30.02} = 6$$



5) Sample (3.585g) contains 1.388g of C, 0.345g of H, 1.850g O and its molar mass is 62g. What is molecular formula of this substance?

$$\text{C} \quad \frac{1.388}{12.01} = 38.7\% \rightarrow \frac{38.7\%}{12.01} = \frac{3.22}{12.01} = 1$$

$$\text{CH}_3\text{O} = \text{EFM} (12 + 3 + 16) = 62$$

$$\text{H} \quad \frac{0.345}{1.01} = \frac{9.12\%}{1.01} = \frac{9.52}{3.22} = 3$$

$$\frac{62}{31.03} = 2$$

$$\text{O} \quad \frac{1.850}{16.00} = \frac{51.6\%}{16.00} = \frac{3.22}{3.22} = 1$$



C_2H_4	C_2H_6	C_2H_8	C_2H_{10}
32	30	36	38
64	60	72	76
96	90	108	112

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Table of Contents 33 Physical Science Answers Formula Worksheet 3

1. Understanding the eBook 33 Physical Science Answers Formula Worksheet 3
 - The Rise of Digital Reading 33 Physical Science Answers Formula Worksheet 3
 - Advantages of eBooks Over Traditional Books
2. Identifying 33 Physical Science Answers Formula Worksheet 3
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 33 Physical Science Answers Formula Worksheet 3
 - User-Friendly Interface
4. Exploring eBook Recommendations from 33 Physical Science Answers Formula Worksheet 3
 - Personalized Recommendations

- 33 Physical Science Answers Formula Worksheet 3 User Reviews and Ratings
- 33 Physical Science Answers Formula Worksheet 3 and Bestseller Lists

5. Accessing 33 Physical Science Answers Formula Worksheet 3 Free and Paid eBooks

- 33 Physical Science Answers Formula Worksheet 3 Public Domain eBooks
- 33 Physical Science Answers Formula Worksheet 3 eBook Subscription Services
- 33 Physical Science Answers Formula Worksheet 3 Budget-Friendly Options

6. Navigating 33 Physical Science Answers Formula Worksheet 3 eBook Formats

- ePUB, PDF, MOBI, and More
- 33 Physical Science Answers Formula Worksheet 3 Compatibility with Devices
- 33 Physical Science Answers Formula Worksheet 3 Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 33 Physical Science Answers Formula Worksheet 3
- Highlighting and Note-Taking 33 Physical Science Answers Formula Worksheet 3
- Interactive Elements 33 Physical Science Answers Formula Worksheet 3

8. Staying Engaged with 33 Physical Science Answers Formula Worksheet 3

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 33 Physical Science Answers Formula Worksheet 3

9. Balancing eBooks and Physical Books 33 Physical Science Answers Formula Worksheet 3

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 33 Physical Science Answers Formula Worksheet 3

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 33 Physical Science Answers Formula Worksheet 3

- Setting Reading Goals 33 Physical Science Answers Formula Worksheet 3
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 33 Physical Science Answers Formula Worksheet 3

- Fact-Checking eBook Content of 33 Physical Science Answers Formula Worksheet 3

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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