

PART B

ESSAY

[100 marks]

2 hours

Answer four questions in all: three questions from Section I and one question from either Section II or Section III.

All questions carry equal marks.

Credit will be given for clarity of expression and orderly presentation of material.

SECTION I

FOR ALL CANDIDATES

Answer three questions from this section.

I. (a) Define each of the following terms and indicate one use of each:
 (i) Nuclear fission;
 (ii) Nuclear fusion. [4 marks]

(b) Alpha particle emission by $^{235}_{92}\text{U}$ produces an element A. Beta particle emission by the particle A produces another element B. Element B also undergoes alpha particle emission to produce $^{237}_{90}\text{Ac}$. Write balanced equations to represent the above statement. [4 marks]

(c) The models below represent the filling of orbitals in an atom.



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II

State which rule(s) is/are violated or obeyed by each model. [6 marks]

(d) Explain why the boiling point of H_2S with relative molecular mass of 34 is lower than that of H_2O with relative molecular mass of 18. [4 marks]

(e) HCl is passed into each of the following solvents:

(i) water;
 (ii) methylbenzene.

I. State the effect of each solution on blue litmus paper. (Turn to red for I)
 II. Compare the electrical conductivities of the two solutions. [4 marks]

(f) Zinc dust is added to copper (II) tetracoxosulphate (VI) solution.

State (i) what is observed;

(ii) the type of reaction that occurs. [2 marks]

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Umesh Ramnarain

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