

Chapter 7 Chemical Reactions

Summary

7.1 Describing Reactions

- In a chemical reaction, the substances that undergo change are called **reactants**.
 - The new substances formed as a result of that change are called **products**.
 - A **chemical equation** is a representation of a chemical reaction in which the reactants and products are expressed as formulas.
- ☞ The law of conservation of mass states that mass is neither created nor destroyed in a chemical reaction.
- During a chemical reaction, the mass of the products is always equal to the mass of the reactants.
- ☞ In order to show that mass is conserved during a reaction, a chemical equation must be balanced.
- In a balanced chemical equation, the number of atoms on the left side of the equation equals the number of atoms on the right.
 - You can balance a chemical equation by changing the coefficients.
 - Coefficients are the numbers that appear before the formulas.
- ☞ Because chemical reactions often involve large numbers of small particles, chemists use a counting unit called the mole to measure amounts of a substance.
- A mole (mol) is an amount of a substance that contains approximately 6.02×10^{23} particles of that substance. This number is known as Avogadro's number.
 - The mass of one mole of a substance is called a **molar mass**. For an element, the molar mass is the same as its atomic mass expressed in grams.
- ☞ In chemical reactions, the mass of a reactant or product can be calculated by using a balanced chemical equation and molar masses of the reactants and products.

7.2 Types of Reactions

- ☞ Some general types of chemical reactions are synthesis reactions, decomposition reactions, single-replacement reactions, double-replacement reactions, and combustion reactions.
- A **synthesis reaction** is a reaction in which two or more substances react to form a single substance. The general equation for a synthesis reaction is $A + B \longrightarrow AB$.

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