

**Question 1:**

A solution of a substance 'X' is used for white washing.

(i) Name the substance 'X' and write its formula.

(ii) Write the reaction of the substance 'X' named in (i) above with water.

**Answer 1:**

(i) The substance 'X' is calcium oxide. Its chemical formula is  $\text{CaO}$ .

(ii) Calcium oxide reacts vigorously with water to form calcium hydroxide (slaked lime)



**Question 2:**

Why is the amount of gas collected in one of the test tubes in Activity 1.7 double of the amount collected in the other? Name this gas.

**Answer 2:**

During the Electrolysis of water, hydrogen and oxygen is get separated by the electricity. Water ( $\text{H}_2\text{O}$ ) contains two parts hydrogen and one part oxygen. Since hydrogen goes to one test tube and oxygen goes to another, the amount of gas collected in one of the test tubes is double of the amount collected in the other.

**Question 1:**

Why does the colour of copper sulphate solution change when an iron nail is dipped in it?

**Answer 1:**

When an iron nail is dipped in a copper sulphate solution, iron (which is more reactive than copper) displaces copper from copper sulphate solution forming iron sulphate, which is green in colour.



Therefore, the blue colour of copper sulphate solution fades and green colour appears.

**Question 2:**

Give an example of a double displacement reaction other than the one given in Activity 1.10.

**Answer 2:**

Sodium carbonate reacts with calcium chloride to form calcium carbonate and sodium chloride.



In this reaction, sodium carbonate and calcium chloride exchange ions to form two new compounds. Hence, it is a double displacement reaction.

**Question 3:**

Identify the substances that are oxidised and the substances that are reduced in the following reactions.



**Answer 3:**

(i) Sodium (Na) is oxidised as it gains oxygen and oxygen gets reduced.

(ii) Copper oxide (CuO) is reduced to copper (Cu) while hydrogen ( $\text{H}_2$ ) gets oxidised to water ( $\text{H}_2\text{O}$ ).

# Answers To Mastering Chemistry Ch 1 Questions

**Catherine E. Housecroft, Edwin C.  
Constable**

## Answers To Mastering Chemistry Ch 1 Questions:

CBSE Class 12 Mastering Organic Reactions MCQ With Answers Priti Singhal,2024-11-11 The primary objective of this book is to serve as a comprehensive guide for students educators and researchers by focusing on reaction mechanisms practical applications and problem solving techniques Organic chemistry is not just about memorizing equations and formulas it is about understanding how molecules interact change and influence each other under different conditions With that in mind this book emphasizes the logic and patterns behind organic reactions making it easier for readers to apply concepts across a variety of scenarios Each chapter of this book builds upon foundational knowledge ensuring a progressive learning experience From nucleophilic substitutions to pericyclic reactions and from oxidation reduction mechanisms to named reactions we cover both fundamental and advanced topics to cater to students at all levels Real world examples have been integrated throughout the chapters to show how organic reactions play essential roles in pharmaceuticals biochemistry agriculture and environmental science This approach bridges the gap between theory and practical applications helping readers appreciate the relevance of organic chemistry in daily life

**Chemistry, Student Study Guide** James E. Brady,Fred Senese,2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH<sub>4</sub> from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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