

Chemthink: CHEMICAL REACTIONS

- (a) Go to **Chemthink**. Log in using your user name and password. (If you forgot, you need to **WAIT QUIETLY** until your teacher is done introducing the lesson.)
- (b) Scroll to **CHEMICAL REACTIONS** then the tutorial under the Chemical Reactions menu.
- (c) Complete the tutorial. As you complete the tutorial, write down the answers for the questions below.
- (d) Once you have successfully completed the tutorial and the worksheet, complete the **Question Set**.

CHEMICAL REACTIONS TUTORIAL QUESTIONS:

1. Starting materials in a chemical reaction are called _____.
2. The ending materials in a chemical reaction are called _____.
3. The arrow indicates a _____ has taken place.
4. All reactions have one thing in common: there is a _____ of chemical bonds.
5. Chemical reactions always involve _____ old bonds, _____ new bonds, or both.
6. In all reactions we will have all of the _____ at the end that we had at the start.
7. In every reaction there can never be any _____ atoms or _____.
8. Chemical reactions only _____ in the atoms that are already there.
9. Let's represent a reaction on paper. For example: hydrogen gas (H_2) reacts with oxygen gas (O_2) to form water (H_2O).



If we use only the atoms shown, we'd have _____ atoms of H and _____ atoms of O in reactants. This would make _____ molecules of H_2O , but we'd have _____ atoms of O leftover. However, this reaction only makes H_2O .

Remember: reactions are kept balanced to 1 molecule each of reactants. We can have as many as we need to balance the chemical equation.

Balancing chemical reactions sheet:

- (a) What atoms are present before (in the reactants) and after (in the products)
- (b) What mass of each reactant and product is present before and after.
- (c) So to make H_2O from oxygen gas and hydrogen gas, the balanced equation would be:



# of atoms in Reactants	Element	# of atoms in Products
	H	
	O	

- (d) This idea is called the _____.
- (e) There must be the same _____ and the same number of _____ before the reaction (in the reactants) and after the reaction (in the products).
- (f) What is the balanced equation for this reaction? _____ Ca + _____ Cl_2 \rightarrow _____ $CaCl_2$

Answers To Chemthink Questions Tutorial Test

H. Stephen Stoker,Danny V. White

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