

14.2 Classifying Reactions

READ

Chemical reactions may be classified into different groups according to the reactants and products. The five major groups of chemical reactions are summarized below.

Synthesis reactions - when two or more substances combine to form a new compound.

- *General equation:* $A + B \rightarrow AB$
- *Example:* When rust forms, iron reacts with oxygen to form iron oxide (rust).
 $4\text{Fe (s)} + 3\text{O}_2 \text{ (g)} \rightarrow 2\text{Fe}_2\text{O}_3 \text{ (s)}$

Decomposition reactions - when a single compound is broken down to produce two or more smaller compounds.

- *General equation:* $AB \rightarrow A + B$
- *Example:* Water can be broken down into hydrogen and oxygen gases.
 $2\text{H}_2\text{O (l)} \rightarrow 2\text{H}_2 \text{ (g)} + \text{O}_2 \text{ (g)}$

Single displacement reactions - when one element replaces a similar element in a compound.

- *General equation:* $A + BX \rightarrow AX + B$
- *Example:* When iron is added to a solution of copper chloride, iron replaces copper in the solution and copper falls out of the solution.
 $\text{Fe (s)} + \text{CuCl}_2 \text{ (aq)} \rightarrow \text{Cu (s)} + \text{FeCl}_2 \text{ (aq)}$

Double displacement reactions - when ions from two compounds in solution exchange places to produce two new compounds.

- *General equation:* $AX + BY \rightarrow AY + BX$
- *Example:* When carbon dioxide gas is bubbled into lime water, a precipitate of calcium carbonate is formed along with water.
 $\text{CO}_2 \text{ (g)} + \text{Ca(OH)}_2 \text{ (aq)} \rightarrow \text{CaCO}_3 \text{ (s)} + \text{H}_2\text{O (l)}$

Combustion reactions - when a carbon compound reacts with oxygen gas to produce carbon dioxide and water vapor. Energy is released from the reaction.

- *General equation:* $\text{Carbon Compound} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O} + \text{energy}$
- *Example:* The combustion of methane gas.
 $\text{CH}_4 \text{ (g)} + 2\text{O}_2 \rightarrow \text{CO}_2 \text{ (g)} + 2\text{H}_2\text{O (g)}$

EXAMPLE

Classify the following reaction as synthesis, decomposition, single displacement, double displacement, or combustion. Explain your answer.



Answer: Displacement. Magnesium replaces copper in the compound.

Answers To And 21 Classifying Chemical Reactions

Stanley E. Manahan



Answers To And 21 Classifying Chemical Reactions:

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