

Chemistry First Semester Review

Key

Introduction to Chemistry

Terms to Define

- qualitative data
- quantitative data
- independent variable
- dependent variable
- control
- constants
- scientific theory
- scientific law
- standard of measurement
- base unit
- derived unit
- accuracy
- precision

Identify the following examples of data as qualitative (Q/L) or quantitative (Q/T).

Q/L 1. the solution is blue
 Q/L 2. the metal has luster
 Q/L 3. the mass is 10.5 g
 Q/T 4. the temperature is 12 °C
 Q/L 5. the length is 5.0 mm
 Q/L 6. the liquid conducts electricity

Identify the SI base unit used for each of the following:

time	s
length	m
mass	kg
temperature	K
amount of a substance	mol

Metric Conversions

convert	to	
0.0032 mg	to	3.2×10^{-6}
2,390 km	to	2.39×10^3
0.009 g	to	9×10^{-6}
0.003 Mg	to	3.3×10^{-4}
0.00052 s	to	0.52
3.42 µg	to	3.42×10^{-6}

Temperature Conversions

Convert the following:

1) 362°C = 635 K
 2) 25°C = 298 K
 3) 300 K = 27 °C
 4) 0 K = -273 °C
 5) 25 K = -248 °C

You design an experiment to determine which household cleaner best prevents the growth of bacteria. You gather five different household cleaners and some petri dishes with bacteria samples. Determine each of the following for your experiment:

independent variable: type of cleaner
 dependent variable: how well cleaners control bacteria
 control: petri dish w/ bacteria, no cleaner
 constants: temp, strain of bacteria, sunlight, etc.

Scientific Notation

Convert the following to scientific or standard notation:

1) $8,320,000$ = 8.32×10^6
 2) 0.000937 = 9.37×10^{-4}
 3) $7,500,000,000$ = 7.5×10^9
 4) 5.22×10^4 = $52,200$
 5) 9.214×10^8 = $9,214,000,000$
 6) 8.5×10^{-6} = 0.0000085

Scientific Notation

Perform the following calculations:

1) $(2.1 \times 10^4) \times (8.33 \times 10^4)$ = 1.7×10^{10}
 2) $(7.50 \times 10^{10}) / (4.212 \times 10^9)$ = 1.78×10^1
 3) $(4.4 \times 10^9) + (7.2 \times 10^9)$ = 5.1×10^{10}
 4) $(5.29 \times 10^6) - (6.04 \times 10^5)$ = 5.51×10^5

Answers To 2nd First Semester Chemistry Review

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Answers To 2nd First Semester Chemistry Review:

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