

Chemistry First Semester Review

Introduction to Chemistry

Terms to Define

- qualitative data
- quantitative data
- independent variable
- dependent variable
- control
- constants
- scientific theory
- scientific law
- standard of measurement
- base unit
- derived unit
- accuracy
- precision

Identify the following examples of data as qualitative (QL) or quantitative (QT).

- QL 1. the solution is blue
 QL 2. the metal has luster
 QN 3. the mass is 105 g
 QN 4. the temperature is 12 °C
 QN 5. the length is 5.0 mm
 QL 6. the liquid conducts electricity

Identify the SI base unit used for each of the following:

time	s
length	m
mass	kg
temperature	K
amount of a substance	mol

Metric Conversions

convert	to	
0.0032 mg	g	3.2×10^{-6}
2390 km	m	2.39×10^6
0.009 g	mg	9×10^4
0.033 Mg	cg	3.3×10^4
0.00032 s	ms	0.32
3.42 µg	kg	3.42×10^{-9}

Temperature Conversions

Convert the following:

- 1) $362\text{ °C} = 635\text{ K}$
 2) $25\text{ °C} = 298\text{ K}$
 3) $300\text{ K} = 27\text{ °C}$
 4) $0\text{ K} = -273\text{ °C}$
 5) $25\text{ K} = -248\text{ °C}$

Experimental Design

You design an experiment to determine which household cleaner best prevents the growth of bacteria. You gather five different household cleaners and some petri dishes with bacteria samples. Determine each of the following for your experiment:

independent variable: type of cleaner
 dependent variable: how well cleaners control bacteria
 control: petri dish w/ bacteria, no cleaner
 constants: temp, strain of bacteria, sunlight, etc

Scientific Notation

Convert the following to scientific or standard notation:

- 1) $8,320,000 = 8.32 \times 10^6$
 2) $0.000937 = 9.37 \times 10^{-4}$
 3) $7,500,000,000 = 7.5 \times 10^9$
 4) $5.22 \times 10^4 = 52,200$
 5) $9.214 \times 10^6 = 9,214,000$
 6) $8.5 \times 10^{-8} = 0.00000085$

Scientific Notation

Perform the following calculations:

- 1) $(2.1 \times 10^5) \times (8.33 \times 10^4) = 1.7 \times 10^{10}$
 2) $(7.50 \times 10^{30}) / (4.212 \times 10^{30}) = 1.78 \times 10^{-3}$
 3) $(4.4 \times 10^6) + (7.2 \times 10^6) = 5.1 \times 10^6$
 4) $(5.29 \times 10^9) - (6.04 \times 10^9) = -5.51 \times 10^9$

Answers To 2nd First Semester Chemistry Review

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Answers To 2nd First Semester Chemistry Review:

Chemistry (Teacher Guide) Dr. Dennis Englin, 2018-02-26 This book was created to help teachers as they instruct students through the Master's Class Chemistry course by Master Books. The teacher is one who guides students through the subject matter, helps each student stay on schedule and be organized, and is their source of accountability along the way. With that in mind, this guide provides additional help through the laboratory exercises as well as lessons, quizzes, and examinations that are provided along with the answers. The lessons in this study emphasize working through procedures and problem solving by learning patterns. The vocabulary is kept at the essential level. Practice exercises are given with their answers so that the patterns can be used in problem solving. These lessons and laboratory exercises are the result of over 30 years of teaching home school high school students and then working with them as they proceed through college. Guided labs are provided to enhance instruction of weekly lessons. There are many principles and truths given to us in Scripture by the God that created the universe and all of the laws by which it functions. It is important to see the hand of God and His principles and wisdom as it plays out in chemistry. This course integrates what God has told us in the context of this study.

Features
Each suggested weekly schedule has five easy-to-manage lessons that combine reading and worksheets. Worksheets, quizzes, and tests are perforated, and three-hole punched materials are easy to tear out, hand out, grade, and store. Adjust the schedule and materials needed to best work within your educational program. Space is given for assignments, dates. There is flexibility in scheduling. Adapt the days to your school schedule.

Workflow
Students will read the pages in their book and then complete each section of the teacher guide. They should be encouraged to complete as many of the activities and projects as possible as well. Tests are given at regular intervals with space to record each grade.

About the Author
DR DENNIS ENGLIN earned his bachelor's from Westmont College, his master of science from California State University, and his EdD from the University of Southern California. He enjoys teaching animal biology, vertebrate biology, wildlife biology, organismic biology, and astronomy at The Master's University. His professional memberships include the Creation Research Society, the American Fisheries Association, Southern California Academy of Sciences, Yellowstone Association, and Au Sable Institute of Environmental Studies.

Queries, 1886 **Russian Chemical Reviews**, 2003 Practical Druggist and Pharmaceutical Review of Reviews Benjamin Lillard, 1922 *Bulletin of the Chemical Society*, 1973 *The Lancet*, 1881 **British Medical Journal**, 1882 *Saturday Review*, 1887 **E3 Chemistry Guided Study Book - 2018 Home Edition (Answer Key Included)**

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Table of Contents Answers To 2nd First Semester Chemistry Review

1. Understanding the eBook Answers To 2nd First Semester Chemistry Review
 - The Rise of Digital Reading Answers To 2nd First Semester Chemistry Review
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers To 2nd First Semester Chemistry Review
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers To 2nd First Semester Chemistry Review
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers To 2nd First Semester Chemistry Review

- Personalized Recommendations
 - Answers To 2nd First Semester Chemistry Review User Reviews and Ratings
 - Answers To 2nd First Semester Chemistry Review and Bestseller Lists
5. Accessing Answers To 2nd First Semester Chemistry Review Free and Paid eBooks
 - Answers To 2nd First Semester Chemistry Review Public Domain eBooks
 - Answers To 2nd First Semester Chemistry Review eBook Subscription Services
 - Answers To 2nd First Semester Chemistry Review Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Answers To 2nd First Semester Chemistry Review Compatibility with Devices
 - Answers To 2nd First Semester Chemistry Review Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers To 2nd First Semester Chemistry Review
 - Highlighting and Note-Taking Answers To 2nd First Semester Chemistry Review
 - Interactive Elements Answers To 2nd First Semester Chemistry Review
 8. Staying Engaged with Answers To 2nd First Semester Chemistry Review
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers To 2nd First Semester Chemistry Review
 9. Balancing eBooks and Physical Books Answers To 2nd First Semester Chemistry Review
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers To 2nd First Semester Chemistry Review
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Answers To 2nd First Semester Chemistry Review
 - Setting Reading Goals Answers To 2nd First Semester Chemistry Review
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Answers To 2nd First Semester Chemistry Review

- Fact-Checking eBook Content of Answers To 2nd First Semester Chemistry Review
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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