

Approach to Arterial Blood Gases (ABGs)

1. Adequate oxygenation?

Normal ABG on Room Air = 7.40 / 40 / 90 / 24

Normal pCO_2 should be $\text{pCO}_2 \approx 4-5$ (pCO_2 in room air = 30)

$\text{pCO}_2 < 5$, then expected

Possible hypoxemia

2. Define the acid-base disturbance

$\text{pH} < 7.40 =$ Acidosis

$\text{pH} > 7.40 =$ Alkalosis

See slide on
Hypoxemia, Pathogenesis and Clinical Findings

3. Identify the primary (Δ^+) process

4. Compensation mechanisms in lungs and kidneys attempt to keep/retain CO_2 or HCO_3^- to maintain normal pH. Determination of compensation is appropriate (appropriate normal ratio of change in pCO_2 : HCO_3^- is shown in green boxes). If inappropriate, there is a secondary (Δ^-) process accounting

$\Delta^+ \text{ HCO}_3^-$

$\Delta^+ \text{ HCO}_3^-$

$\Delta^- \text{ HCO}_3^-$

5. Calculate anion gap (AG) to determine the presence and type of metabolic acidosis. This is done regardless of the Δ^+ or Δ^- process, as there may also be a hidden metabolic acidosis

Calculate $\text{AG} = \text{Na}^+ - \text{Cl}^- - \text{HCO}_3^-$ (normal is ≈ 12)

6. HAGMA: refer to normal blood buffer system, acid gain should match bicarbonate loss. If not, identify if another process is causing an inappropriate loss or gain of HCO_3^- .
 $\text{Gain of acid (AG)} = \text{AG} - 12$
 $\text{Loss of } \text{HCO}_3^- (\text{AHCO}_3^-) = 24 - \text{HCO}_3^-$

Classically, the compensating anion is Cl^- , maintaining a normal AG

Where the cause is addition of an abnormal acid, HCO_3^- is used up as a buffer, and the compensating anion is the abnormal acid's conjugate base - which is not measured in the AG calculation

7. Calculate Cation Gap (net positive) or Urine net charge (for HAGMA) to narrow the etiology

$\text{AG} < 12$

$\text{AG} > 12$

If no metabolic acidosis was identified in steps 2-4:

If a Δ^+ or Δ^- metabolic acidosis was identified in steps 2-4:

High Anion Gap Metabolic Acidosis (HAGMA)

No metabolic acidosis present

Normal Anion Gap Metabolic Acidosis (NAGMA)

Calculate $\text{AG} = \text{Na}^+ - \text{Cl}^- - \text{HCO}_3^-$

In this scenario, a normal total cation should be exceeding excess acid

$\text{Excess acid is normally excreted as ammonium } (\text{NH}_4^+) \text{ with } \text{Cl}^-$

$\text{Na}^+ - \text{Cl}^- = \text{AG}$

The more Cl^- in the urine relative to cations, the more acid is being excreted

$\text{Calculate urine } (\text{U}) \text{ net charge to determine cause of NAGMA: } \text{U}_{\text{Cl}^-} - \text{U}_{\text{Na}^+} - \text{U}_{\text{K}^+}$

$\text{HAGMA} = \text{Metabolic acidosis}$

Use U net charge

$\text{HAGMA} = \text{Metabolic acidosis}$

$\text{HAGMA} = \text{Metabolic acidosis}$

Use U net charge

$\text{HAGMA} = \text{Metabolic acidosis}$

$\text{HAGMA} = \text{Metabolic acidosis}$

Appropriately large Cl^- excretion $\Delta^+ \text{ HCO}_3^-$ loss is the issue

$\text{HAGMA} = \text{Metabolic acidosis}$

$\text{HAGMA} = \text{Metabolic acidosis}$

Inappropriately small Cl^- excretion $\Delta^- \text{ HCO}_3^-$ gain is the issue

$\text{HAGMA} = \text{Metabolic acidosis}$

$\text{HAGMA} = \text{Metabolic acidosis}$

See slide on Normal Anion Gap Metabolic Acidosis Pathogenesis and Laboratory Findings

$\text{HAGMA} = \text{Metabolic acidosis}$

$\text{HAGMA} = \text{Metabolic acidosis}$

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* 10th or later of publication

$\text{HAGMA} = \text{Metabolic acidosis}$

$\text{HAGMA} = \text{Metabolic acidosis}$

Legend:

Diagnose (Rate)

Mechanism

Lab Finding/Calculation

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See slide on Normal Anion Gap Metabolic Acidosis Pathogenesis and Laboratory Findings

See slide on High Anion Gap Metabolic Acidosis Pathogenesis and Laboratory Findings

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F. A. Buttress, Henry J. Heaney

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