

QUESTION 1

1. Choose a description from column A and column B that matches a word in column A, write only the letter (A-J) next to the question number.

COLUMN A	COLUMN B
1.1 Conflict	A. Union of interest purpose within the group.
1.2 Solidarity	B. Managers have relatively few subordinates reporting to them.
1.3 Narrow span of control	C. Clash between two people.
1.4 Vision	D. Broad guide on what the enterprise is doing.
1.5 Policy	E. Roadmap of what the organisational future makes up will be.

(25)

Question 2

2. 1 Use the information provided to draw up organogram of Northern FET College.

NORTHERN FET COLLEGE

JOB TITLE	NUMBER OF EMPLOYEES
Marketing manager	1
CEO	1
Principal	1
HOD Engineering and Business Lectures.	2
OB Business and OT Engineering Administration manager	10
Finance manager	1

(10)

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RC Schank

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web k sp of calcium hydroxide 1 author j m mccormick last update october 13 2013 introduction an equilibrium constant k is related to ΔG for a process through eqn 1 and as long as ΔH and ΔS are independent of temperature eqn 1 can be combined with eqn 2 to determine ΔH and ΔS from the temperature dependence of k

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide ca oh 2 by measuring the concentration of ca oh 2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of ca oh is represented in equation form as shown below $\text{Ca}^{2+} + 2\text{OH}^- \rightleftharpoons \text{Ca}(\text{OH})_2$ the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a ph regulating agent notice that calcium hydroxide is divalent and thus has twice

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $Ca(OH)_2 \rightleftharpoons Ca^{2+} + 2OH^-$ 5 x 10 2 m calcium ions 5 x 10 2 m 2 0 m hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is $Ca(OH)_2 \rightleftharpoons Ca^{2+} + 2OH^-$

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web the chemical equation used was $Ca(OH)_2 + 2HCl \rightarrow CaCl_2 + 2H_2O$ the solubility product was determined to be $K_{sp} = Ca^{2+} \cdot (OH^-)^2$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as $Ca_3(PO_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{sp} = Ca^{2+} \cdot (PO_4^{3-})^2$

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web lab report 8 determinin of calcium hydroxide the purpose of this was to titrat e a saturated $Ca(OH)_2$ soultion with a standard hcl solution and find the ksp or the oh concentration was f ound and then the concentration was found and u for this experiment a titration of calcium hydroxide with hydrochlo

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for this experiment the titration of calcium

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web question chem 182 experiment 8 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water solution of Ca(OH)_2 is represented in equation form lution of Ca^{2+} and OH^- ionic solid that is sparingly soluble in water a saturated aqueous as shown below

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for Ca(OH)_2 was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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web we will filter the solid calcium hydroxide from the solution and then titrate the dissolved calcium hydroxide in solution with a known concentration of hydrochloric acid using titrimetric calculations we will determine the Ca^{2+} and OH^- concentrations in solution

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web calculate the ksp for calcium hydroxide $\text{K}_{\text{sp}} = 5 \times 10^{-6}$ find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error accepted value for ksp for sodium hydroxide is 5×10^{-5}

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