

Electron Configuration (Ions / Valence) - Worksheet #6

Symbol	# of Electrons	Electron Configuration (Use EITHER method, Core or "Long")
Ca^{+2}	18	$Ca^{+2} : [Ar]^{18}$
Se^{-2}	32	$Se^{-2} : [Ar]^{18} 3s^2 3p^6$
Sn^{+2}	22	$Sn^{+2} : [Kr]^{18} 5s^2 4s^2$
Zn^{+2}	18	$Zn^{+2} : [Ar]^{18} 3d^{10}$
Sr^{+2}	22	$Sr^{+2} : [Kr]^{18}$
O^{-2}	6	$O^{-2} : [He]^{2} 2s^2 2p^4$
Cs^{+1}		$Cs^{+1} : [Xe]^{18}$
I^{-}	8	$I^{-} : [Kr]^{18} 5s^2 4s^2 3p^6$
Fe^{+3}	7	$Fe^{+3} : [Ar]^{18} 4s^1 3d^5$
Cu^{+2}	9	$Cu^{+2} : [Ar]^{18} 4s^2 3d^9$

Directions: State how many valence electrons

1. O _____ 6 _____

2. P _____ 5 _____

3. V _____

4. Ca _____ 2 _____

5. Xe _____ 0 _____

6. Hg _____

7. Te _____ 6 _____

8. Cl⁻ _____

9. I⁻³ _____

10. Zn⁺² _____

11. Ge⁻⁴ _____

12. Sb⁺³ _____

13. O⁻ _____

14. Tc⁻⁴ _____

5 2 Electron Configurations Answer Key

IM Harris

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