

Electron Configuration (Ions / Valence) - Worksheet #6

Symbol	# of Electrons	Electron Configuration (Use EITHER method. Core or "Long")
Ca^{+2}	18	$\text{Ca}^{+2} : [\text{Ar}]^{18}$
Se^{-2}	34	$\text{Se}^{-2} : [\text{Ar}]^{18} 4s^2 3d^{10} 4p^6$
Sn^{+2}	50	$\text{Sn}^{+2} : [\text{Kr}]^{36} 5s^2 4d^{10}$
Zn^{+2}	28	$\text{Zn}^{+2} : [\text{Ar}]^{18} 3d^{10}$
Sr^{+2}	36	$\text{Sr}^{+2} : [\text{Kr}]^{36}$
O^{-2}	10	$\text{O}^{-2} : [\text{He}]^2 2s^2 2p^6$
Cs^{+1}	54	$\text{Cs}^{+1} : [\text{Xe}]^{54}$
I^{-1}	54	$\text{I}^{-1} : [\text{Xe}]^{54} 5s^2 4d^{10} 5p^6$
Fe^{+3}	23	$\text{Fe}^{+3} : [\text{Ar}]^{18} 4s^1 3d^5$
Cu^{+2}	27	$\text{Cu}^{+2} : [\text{Ar}]^{18} 4s^2 3d^9$

Directions: State how many valence electrons

- O 6
- P 5
- V 5
- Ca 2
- Xe 0
- Hg 0
- Te 6

- Cl 7
- I^{-1} 7
- Zn^{+2} 0
- Ge^{+4} 0
- Sb^{+3} 0
- O^{-2} 0
- Tc^{+4} 0

5 2 Electron Configurations Answer Key

IM Harris



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