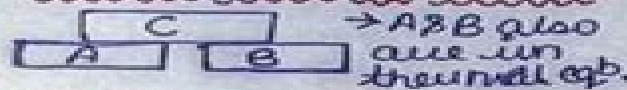


Branch of science with conversion of heat and work.

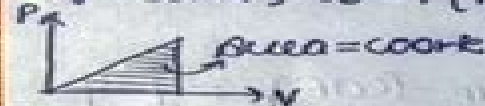
* Zeroth Law of Thermodynamics



* First Law of Thermodynamics

$$\{dQ = dW + \Delta U\}$$

- work done $\Rightarrow dW = \int P dv$
- if P const; $W = P(V_2 - V_1)$



- $\Delta U = \frac{f}{2} n R \Delta T$
- $\{\Delta U = n C_v \Delta T\}$
- $\{\Delta U = \frac{n R \Delta T}{\gamma - 1}\}$ $\frac{f}{2} = \frac{1}{\gamma - 1}$
- $\{\Delta U = \frac{P_2 V_2 - P_1 V_1}{\gamma - 1}\}$

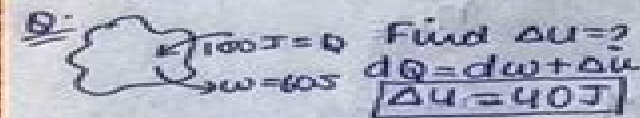
* Sign Conventions

Heat in, work out

$$\Delta U = \frac{f}{2} n R \Delta T$$

$$T \uparrow \rightarrow \Delta T \uparrow \therefore \Delta U \Rightarrow +ve$$

$$T \downarrow \rightarrow \Delta T (-ve) \therefore \Delta U = -ve$$

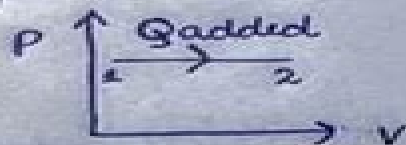


$$Q_{out} = -ve \quad Q_{in} = +ve$$

(i) Isoobaric: ($P=C$)

$$P V = n R T \rightarrow V \propto T$$

$$V/T = \text{constant}$$



HERMODYNAMICS

$$W = \int P \cdot dv \Rightarrow W = P(V_2 - V_1)$$

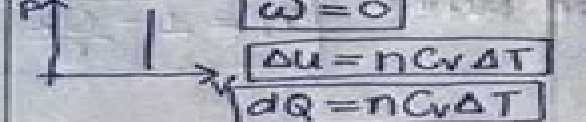
$$\Delta U = n C_v \Delta T \quad \text{for } \Delta U \text{ pure}$$

$$dQ = n C_p \Delta T$$

(ii) Isochoric: vol const

$$P \propto T \rightarrow P/T = \text{constant}$$

Gay Lussac law.



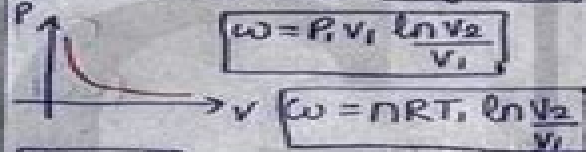
$$W = 0$$

$$\Delta U = n C_v \Delta T$$

$$dQ = n C_v \Delta T$$

(iii) Isothermal (Boyle's law)

$$P V = \text{constant}, \quad x y = \text{const}$$



$$W = P_1 V_1 \ln \frac{V_2}{V_1}$$

$$W = n R T \ln \frac{V_2}{V_1}$$

$$\Delta U = 0 \quad dQ = dW$$

(iv) Adiabatic $Q=0$

Boiling of tyres

$$P V^\gamma = \text{const}$$

$$Q = 0$$

$$dW = -\Delta U$$

$$dW = n R (T_1 - T_2) / (\gamma - 1)$$

$$\Delta U = n C_v \Delta T$$

$$\frac{P_2}{P_1} = \left(\frac{V_1}{V_2}\right)^\gamma, \quad \frac{T_2}{T_1} = \left(\frac{P_2}{P_1}\right)^{\frac{\gamma-1}{\gamma}}$$

$$\frac{T_2}{T_1} = \left(\frac{V_1}{V_2}\right)^{\gamma-1}$$

$$\frac{T_2}{T_1} = \left(\frac{V_1}{V_2}\right)^{\gamma-1}$$

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$$\frac{T_2}{T_1} = \left(\frac{V_1}{V_2}\right)^{\gamma-1}$$

$$\frac{T_2}{T_1} = \left(\frac{V_1}{V_2}\right)^{\gamma-1}$$

(v) Polytropic:

$$P V^n = \text{const}$$

$$Q \neq 0$$

$$dQ = dW + \Delta U$$

$$dW = \frac{n R (T_1 - T_2)}{\gamma - 1}$$

$$\Delta U = \frac{n R \Delta T}{\gamma - 1}$$

$$\text{Polytropic Index}$$

$$\Delta U = n C_v \Delta T$$

$$P = \text{const at } n = 0$$

$$P V = \text{const at } n = 1$$

$$P V^\gamma = \text{const at } n = \gamma$$

$$V = \text{const at } n = \infty$$

$$V = \text{const at } n = \infty$$

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* Cyclic Process:

events:

area enclosed

$\Delta U = P_{in} - P_{out}$

$W = P_{in} - P_{out}$

$Q = P_{in} - P_{out}$

$Q = P_{in} - P_{out}$

$Q = P_{in} - P_{out}$

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