

11 Chemical Reactions

Big idea REACTIONS

11.1 Describing Chemical Reactions

Essential Understanding Chemical reactions are represented by balanced chemical equations.

Lesson Summary

Introduction to Chemical Equations A chemical equation uses symbols, and sometimes words, to show the reactants and products of a chemical reaction.

- ▶ A skeleton equation uses chemical formulas to represent reactants and products, but it does not indicate the relative amounts of each.
- ▶ Anything that enters into a reaction, such as heat or a catalyst, but is not a reactant or product is shown above or below the yields arrow in the equation.

Balancing Chemical Equations Chemical equations are balanced to show that mass is conserved during chemical reactions.

- ▶ A balanced equation shows the relative amounts of reactants and products, and it contains equal numbers of each type of atom on both sides of the equation.
- ▶ Chemical equations are balanced by using coefficients in front of the chemical formulas for the reactants and the products in a skeleton equation.



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Balancing Equations All chemical equations must be balanced because of the law of conservation of mass, which states that matter cannot be created or destroyed. So, the number of atoms that you start with at the beginning of the reaction must equal the number of atoms that you end up with.

For example, the reaction $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$ follows the law of conservation of mass because you start with 2 magnesium atoms and 2 oxygen atoms and you end up with 2 magnesium atoms and 2 oxygen atoms. You can think of the \rightarrow as an = sign.



Turn the page to learn more about balancing equations.

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SJ Ball



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Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application **Prentice Hall Chemistry**, 2000 *Human Biology and Health*, 1997 **Prentice Hall Exploring Life Science** Anthea Maton, 1997 Patai's 1992 Guide to the Chemistry of Functional Groups Saul Patai, 1992-04-08 Patai's 1992 Guide to the Chemistry of Functional Groups Saul Patai The Hebrew University of Jerusalem Israel Ever since the publication of the first volume of The Chemistry of Functional Groups in 1964 the Patai series has acted as an essential reference source to many researchers By the end of 1991 the series consisted of 50 titles bound in 73 volumes containing nearly 900 chapters written by over 1250 authors The aim of this Guide as was that of the previous edition is to present sufficient material on each of the published chapters to allow the researcher to decide whether these chapters are relevant and useful for his or her purpose and thus worth pursuing in full For those who are familiar with only selected volumes from the series the cross referencing between complementary and related chapters from different volumes will be invaluable The Guide is fully indexed by both subject and author thus making it an essential reference tool for all organic chemists *Holt Chemistry* Ralph Thomas Myers, 2004 Prentice Hall Science Series, 1994 Prentice-Hall Staff, 1993 Bioconjugate Techniques Greg T. Hermanson, 1996-01-15 Bioconjugate Techniques is the essential guide to the modification and crosslinking of biomolecules for use in research diagnostics and therapeutics It provides highly detailed information on the chemistry reagent systems and practical applications for creating labeled or conjugate molecules It also describes dozens of reactions with details on hundreds of commercially available reagents and the use of these reagents for modifying or crosslinking peptides and proteins sugars and polysaccharides nucleic acids and oligonucleotides lipids and synthetic polymers Armed with this information and the abundant protocols provided readers will form unique complexes that can be used for detecting quantifying and targeting important analytes This book helps readers make high activity antibody enzymes conjugates immunotoxins immunogen complexes liposome conjugates as well as biotinylated molecules avidin or streptavidin conjugates colloidal gold labeled proteins PEG or dextran complexes labeled oligonucleotide probes and fluorescently tagged or radiolabeled molecules This book is the first to thoroughly capture the entire field of bioconjugate chemistry in a single volume Serves as a practical guide to modification and cross linking technology for research diagnostics and therapeutics Provides useful detailed easy to follow step by step protocols Contains easy to read and easy to understand key concepts for making bioconjugates of all types Efficiently covers the chemistry of bioconjugation the major reagents available for modification and cross linking and the application of these reagents to the

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