



Alternating Current

An alternating quantity (current i or voltage V) is one whose magnitude changes continuously with time between zero and a maximum value and whose direction reverses periodically.

Some graphical representation for alternating quantities



Equation of Alternating Quantities (i or V)

When a coil is rotated rapidly in a strong magnetic field, magnetic flux linked with the coil changes. As a result an emf is induced in the coil and induced current flows through the circuit. These voltage and current are known as alternating voltage and current

(1) **Equation :** Alternating current or voltage varying as sine function can be written as

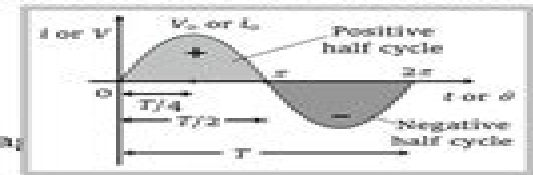
$$i = i_0 \sin \omega t = i_0 \sin 2\pi \nu t = i_0 \sin \frac{2\pi}{T} t$$

$$\text{and } V = V_0 \sin \omega t = V_0 \sin 2\pi \nu t = V_0 \sin \frac{2\pi}{T} t$$

where i and V = Instantaneous values of current and voltage

i_0 and V_0 = Peak values of current and voltage

ω = Angular frequency in rad/sec, ν = Frequency in Hz and T = time period



(2) **About cycle**

(i) The time taken to complete one cycle of variations is called the periodic time or time period.

(ii) Alternating quantity is positive for half the cycle and negative for the rest half. Hence average value of alternating quantity (i or V) over a complete cycle is zero.

(iii) Area under the positive half cycle is equal to area under negative cycle.

(iv) The value of alternating quantity is zero or maximum 2ν times every second. The direction also changes 2ν times every second.

(v) Generally sinusoidal waveform is used as alternating current/voltage.

(vi) At $t = \frac{T}{4}$ from the beginning, i or V reaches to their maximum value.

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Paul O'Brien



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 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l}) \quad \Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + 1/2 \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = 285 \text{ kJ/mole}$

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combustion of magnesium oxide Mg s 1 2O_2 g MgO s we used

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web thus for reaction 9 2 in which 1 mole of H_3O^+ and 1 mole of OH^- combine to form 2 moles of H_2O ΔH 55.8 kJ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}}$

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