

Multiple Choice

1. Which electron configuration corresponds to the ground state of oxygen (O, Z = 8)?

a) $1s^2 2s^2 2p^1 \rightarrow 1s^2 + 2s^2 + 2p^1$
 b) $1s^2 2s^2 2p^2 \rightarrow 1s^2 + 2s^2 + 2p^2$
 c) $1s^2 2s^2 2p^3 \rightarrow 1s^2 + 2s^2 + 2p^3$
 d) $1s^2 2s^2 3s^2 \rightarrow 1s^2 + 2s^2 + 3s^2$

should be 2, not 3

oxygen has 8 electrons in order:
 $1s^2$ $2s^2$ $2p^6$
 $1s^2$ $2s^2$ $2p^2$ $2p^4$
 $1s^2$ $2s^2$ $2p^4$ $3s^2$
 $1s^2$ $2s^2$ $2p^6$ $3s^2$ $3p^2$

$2 \quad 6 \quad 10 \quad 6$
 $2 \quad 6 \quad 10 \quad 6$

$\rightarrow 1s^2 (2s) + 2s^2 (2p) + 2p^6 (3s) + 3p^2 (3p)$
 $= 8 total$

2. According to Hund's rule, how will three electrons occupy the three 2p orbitals?

a) All three electrons pair up in the first 2p orbital.
 b) Two electrons pair in one orbital and one is single in another, spins random.
 c) Each 2p orbital gets one electron with parallel spins before any pairing occurs.
 d) All electrons go into the 2s orbital first.

Hund's Rule = "singlet first"

3. Which statement best describes the Pauli Exclusion Principle?

a) Electrons fill the lowest energy orbitals first.
 b) Orbitals of the same energy each receive single electrons before pairing.
 c) No two electrons in an atom can have the same set of four quantum numbers (i.e., an orbital can hold at most two electrons of opposite spin).
 d) Electrons always spin clockwise in the highest occupied orbital.

this rule explains why orbital can hold 2 electrons and why spin must be opposite

4. Which of these is the correct order for filling orbitals (Aufbau order) around the 3d/4s region?

a) 3d \rightarrow 4s \rightarrow 4p
 b) 4s \rightarrow 3d \rightarrow 4p \rightarrow *not possible: 3d is 3s² and 4s is 1s*
 c) 4p \rightarrow 3d \rightarrow 4s \rightarrow *not possible: 3d is higher than 4s*
 d) 3s \rightarrow 3p \rightarrow 3d

Up each electron occupies the lowest energy orbital available

5. The electron configuration $[Ne]3s^2 3p^1$ corresponds to which neutral element?

a) Sodium (Na)
 b) Magnesium (Mg)
 c) Aluminum (Al)
 d) Silicon (Si)

$1s^2$
 $2s^2$
 $2p^6$
 $3s^2$
 $3p^1$

$10e^- + 2e^- + 6e^- = 18e^-$
 \downarrow
Aluminum
(Al)

*[Ne] indicates 2e- in 1s orbital,
 $+ 2s^2 (2s) + 2p^6 (2p) + 3s^2 (3s) + 3p^1 (3p)$
 \rightarrow Al, atomic number 13*

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Marco Cascella

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