

## Multiple Choice

1. Which electron configuration corresponds to the ground state of oxygen (O,  $Z = 8$ )?

☒ a)  $1s^2 2s^2 2p^4 \rightarrow e^- = 8$   
 Should be 2 not 1 → ~~b)  $1s^2 2s^2 2p^5 \rightarrow e^- = 10$~~   
 c)  $1s^2 2s^2 2p^3 \rightarrow e^- = 8$   
~~d)  $1s^2 2s^2 3s^2 \rightarrow e^- = 6$~~

1s	2s	2p	3s	3p	3d
↑↓	↑↓	↑↑	↑↓	↑↓	↑↓
2	2	4	2	4	6

oxygen has 8 electrons  
fill order

$\rightarrow 1s^2 (2), 2s^2 (2), 2p^4 (4)$   
 $= 8 \text{ total}$

2. According to Hund's rule, how will three electrons occupy the three 2p orbitals?

a) All three electrons pair up in the first 2p orbital.  
☒ b) Two electrons pair in one orbital and one is single in another, spins random.  
 c) Each 2p orbital gets one electron with parallel spins before any pairing occurs.  
 d) All electrons go into the 2s orbital first.

Hund's Rule = "Singles First"

3. Which statement best describes the Pauli Exclusion Principle?

a) Electrons fill the lowest energy orbitals first.  
 b) Orbitals of the same energy each receive single electrons before pairing.  
☒ c) No two electrons in an atom can have the same set of four quantum numbers (i.e., an orbital can hold at most two electrons of opposite spin).  
 d) Electrons always spin clockwise in the highest occupied orbital.

a maximum of 2 electrons  
 can occupy a single orbital  
 with opposite spins

This rule explained why orbitals can hold 2 electrons and why spins must be opposite

4. Which of these is the correct order for filling orbitals (Aufbau order) around the 3d/4s region?

a)  $3d \rightarrow 4s \rightarrow 4p$

☒ b)  $4s \rightarrow 3d \rightarrow 4p$  → although 3d is "3", the 4s is lower in energy and fills before 3d. After that, the electrons continue into 4p.  
 c)  $4p \rightarrow 3d \rightarrow 4s$   
 d)  $3s \rightarrow 3p \rightarrow 3d$

↳ each electron occupies the lowest energy orbital available

5. The electron configuration  $[\text{Ne}]3s^2 3p^1$  corresponds to which neutral element?

a) Sodium (Na)  
 b) Magnesium (Mg)  
☒ c) Aluminum (Al)  
 d) Silicon (Si)



[Ne] accounts for 10 electrons,  
 $+ 3s^2 (2), + 3p^1 (1) = 13 \text{ (total)}$   
 $\rightarrow \text{Al, atomic number 13}$

# 20ap Chem Multiple Choice

**Marco Cascella**



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