

Acid-Base Equilibrium Practice

- What is the molarity of a solution of nitric acid if 0.216g of barium hydroxide is required to neutralize 20.0ml of nitric acid?
- What is the volume of 1.222M sodium hydroxide required to react with:
 - 32.5ml of 0.569M sulfurous acid
 - 5.00g of oxalic acid, $\text{H}_2\text{C}_2\text{O}_4$?
 - 15.0g of concentrated acetic acid?
- What is the volume of 0.885M hydrochloric acid required to react with:
 - 25.00ml of 0.288M aqueous ammonia?
 - 10.0g of sodium hydroxide?
 - 25.0ml of a solution ($d=0.928\text{g}/\text{cm}^3$) containing 10.0% by mass of methylamine, CH_3NH_2 ?
- A capsule of vitamin C, a weak acid, is analyzed by titrating it with a 0.425M sodium hydroxide solution. It is found that 6.20ml of base is required to react with a capsule weighing 0.628g. What is the percentage of vitamin C, $\text{C}_6\text{H}_8\text{O}_6$, in the capsule?
- An artificial beverage contains 12.0g of tartaric acid, $\text{H}_2\text{C}_4\text{H}_4\text{O}_6$, to achieve tartness. It is titrated with a basic solution that has a density of $1.045\text{g}/\text{cm}^3$ and contains 5% by mass, KOH. What volume of the basic solution is required?
- Lactic acid, $\text{C}_3\text{H}_5\text{O}_3$, is the acid in sour milk. A 0.100g sample of pure lactic acid requires 12.95ml of 0.0857M sodium hydroxide for complete reaction. How many moles of hydroxide ions are required to neutralize one mole of lactic acid?
- An aqueous solution contains dissolved $\text{C}_6\text{H}_5\text{NH}_2\text{Cl}$ and $\text{C}_6\text{H}_5\text{NH}_2$. The concentration of $\text{C}_6\text{H}_5\text{NH}_2$ is 0.50M and a pH of 4.20. $K_b = 3.8 \times 10^{-10}$
 - Calculate the concentration of $\text{C}_6\text{H}_5\text{NH}_3^+$ in this buffer solution.
 - Calculate the pH after 4.0grams of HCl is added to 1.0L of solution. Assume no volume change.
- A buffer solution is made up of 355ml each of 0.200M NaHCO_3 and 0.134M Na_2CO_3 , calculate: $K_a = 5.6 \times 10^{-11}$
 - The pH of the buffer
 - The pH of the buffer after the addition of 0.0300mol HCl to 0.710L of buffer solution
 - The pH of the buffer after the addition of 0.0300mol KOH to 0.710L of buffer solution
- Calculate the pH of a solution prepared by mixing 100.0ml of 1.20M ethanolamine, $\text{C}_2\text{H}_5\text{ONH}_2$, with 50.0ml of 1.0M HCl. K_a for $\text{C}_2\text{H}_5\text{ONH}_3^+ = 6.3 \times 10^{-10}$.
- Calculate the pH of the following solution:
 - 2.0M $\text{NaC}_2\text{H}_3\text{O}_2$ / 2.0M $\text{HC}_2\text{H}_3\text{O}_2$
 - The pH of the buffer solution after 50.0ml of 0.15M LiOH has been added to 100.0ml of the buffer solution.

Acid Base Equilibrium Practice Problems Multiple Choice

James C. Hill



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