

Acids and Bases Worksheet

NAME _____

1. Name the following compounds as acids and circle which are weak acids:

- a)
- HNO_3
- b)
- H_2SO_4
- c)
- HF
- d)
- H_2CO_3
- e)
- $\text{HC}_2\text{H}_3\text{O}_2$
- f)
- H_3PO_4

2. Write the formulas for the following bases:

- a) calcium hydroxide b) aluminum hydroxide c) ammonium hydroxide d) lithium hydroxide

3. Complete the following table:

	Acids	Bases
Definition		
Electrolyte in aqueous solution?		
pH		
Taste		
Litmus turns		
Which reacts with metals to form H_2 gas?		

Conjugate acid-base pairs

4. Identify the acid, the base, the conjugate acid and the conjugate base in the reaction:



5. Identify the acid, the base, the conjugate acid and the conjugate base in the reaction:



6. An acid will form the conjugate _____ and a base will form the conjugate _____

Neutralization and Titration.

7. Classify the following acids as monoprotic, diprotic or triprotic:

- a)
- H_2SO_4
- b)
- H_3PO_4
- c)
- HCl
- d)
- $\text{HC}_2\text{H}_3\text{O}_2$
- or
- CH_3COOH
- (same thing)

8. What is the concentration of an unknown NaOH solution if took an average of 15.4mL of 0.100N HCl solution to neutralize 10.0mL of the NaOH solution?9. From the given data, what is the molarity and the normality of the HF solution if the KOH solution was 0.50M?

	Acid, HF	Base, KOH
Initial Volume, mL	5.5	14.9
Final Volume, mL	25.9	24.4

$\text{pH} = -\log[\text{H}_3\text{O}^+]$ To find pH if given $[\text{H}_3\text{O}^+]$, enter the $[\text{H}_3\text{O}^+]$, press Log & then the +/- button
To find $[\text{H}_3\text{O}^+]$ if given pH, enter negative pH, then 2nd function Log

10. A solution has a $[\text{H}_3\text{O}^+]$ of 1×10^{-7} M, what is its pH? ____ Is this an acid, base or is it neutral?

TURN THE PAPER OVER

11. What is the pH of a solution if the $[\text{H}^+] = 7.2 \times 10^{-6}\text{M}$? ____ Is this an acid or base?

Acids To Bases Answer Key

AW Rasmussen



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